

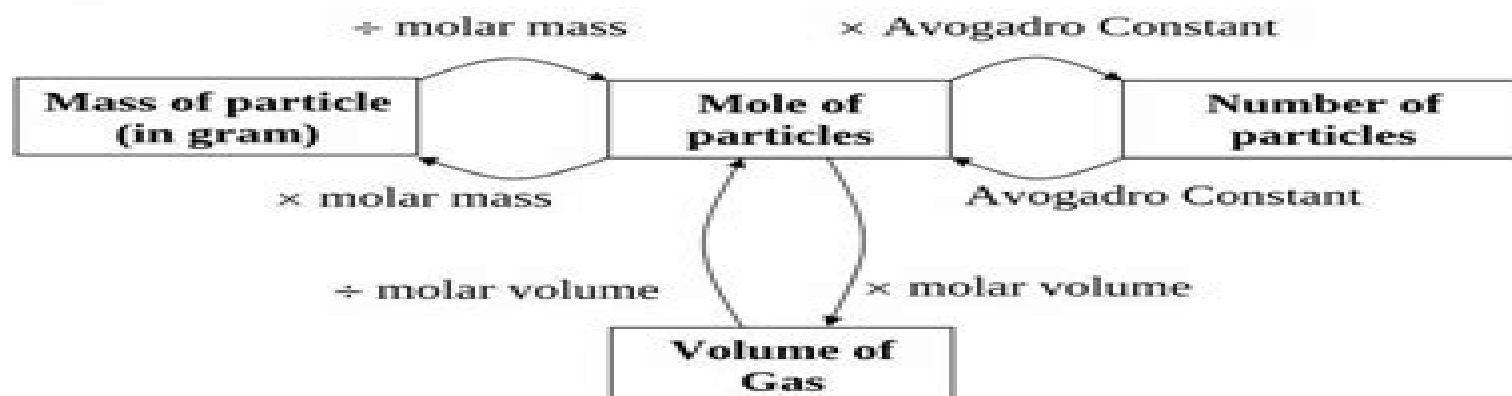
## Short Notes: Form 4 Chemistry

### Chemical Formulae and Equation

#### Calculation

<p><b>For Solid, liquid or gas</b></p> $\text{number of mole} = \frac{\text{mass of substance}}{\text{molar mass}}$ <p>Molar mass = RAM/RMM/RFM in gram</p>	<p><b>For gas (only)</b></p> $\text{number of mole} = \frac{\text{volume of gas}}{\text{molar volume}}$ <p>Molar volume = 24dm<sup>3</sup> at room temperature Molar volume = 22.4dm<sup>3</sup> at s.t.p.</p>
<p><b>For Solution</b></p> $\text{number of mole} = \frac{MV}{1000}$ <p>M = molarity V = Volume of solution in cm<sup>3</sup></p>	<p><b>For quantity of particle(atom,molecule,ion)</b></p> $\text{number of mole} = \frac{\text{quantity of particle}}{6.02 \times 10^{23}}$

#### Summary



# One School Short Notes Form 4 Chemistry

**A Loxley**



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