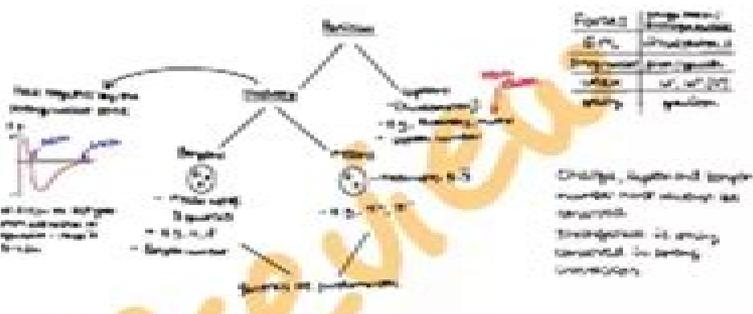
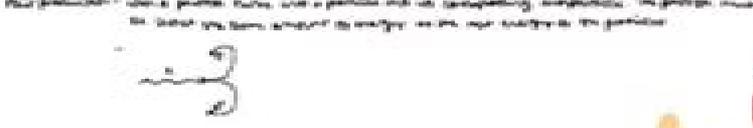


Particles and Quantum



Excitation: when a particle and its corresponding antiparticle collide and their rest energy is converted into photons.

Annihilation: when a particle turns into a particle and its corresponding antiparticle. The photon must have at least the same amount of energy as the rest energy of the particle.



Excitation: Existing more than just excited enough energy to count their atoms.

Alpha particles: Helium nucleus, ${}^4_2\text{He}$

- Highly ionizing, so it is dangerous.
- Relatively penetrating - stopped by a piece of paper or foil on their skin.

Beta: fast moving electron $\rightarrow e^-$

- Medium penetrating and relatively quickly - stopped by a few cm of aluminium.
- Beta decay: ${}^A_Z\text{X} \rightarrow {}^A_{Z+1}\text{Y} + e^- + \bar{\nu}_e$ (antineutrino is almost transparent)



Gamma: high energy EM ray or photon emitted from a high energy nucleus γ

- highly penetrating - stopped by a few m of concrete.

Isotopes: same element - same number of protons

- different relative atomic mass - different number of neutrons

${}^A_Z\text{X}$ where A is mass number, Z is atomic number.

Nuclear charge: $Q = Ze$ (Coulombs)

- charge is mass ratio

Electron rest mass: charge of an electron when accelerated through a pot of one volt

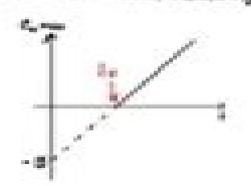
- $1.6 \times 10^{-19} \text{ C}$
- $9.1 \times 10^{-31} \text{ kg}$



Photoelectric effect

Use photons of sufficient energy on absorbed by electrons in the surface of a metal, liberating them.

- $E_{\text{photon}} = hf = \phi + E_{\text{max}}$
- ϕ - work function - minimum energy to liberate electrons
- E_{max} - kinetic energy of electrons
- $E_{\text{photon}} < \phi$ - no electrons emitted
- $E_{\text{photon}} = \phi$ - threshold frequency



The photoelectric effect provides the particle nature of light due to the fact that the maximum kinetic energy of electrons is 0 if the frequency is less than the threshold frequency. This is because the energy of light is quantized into photons. If the energy of a photon is less than the work function, it cannot liberate an electron.

- Stopping potential:** the pd needed to stop any electrons from reaching the other side.
- $E_{\text{photon}} = \phi + eV_s$
 - V_s - stopping potential

De Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle will occur by being described as a packet of waves which propagate through space and time. This is because the electron is described as a wave packet.



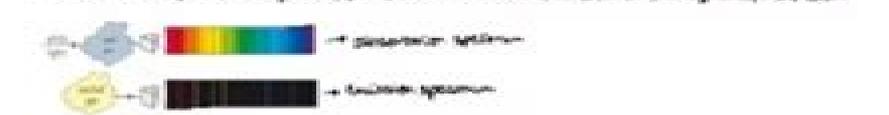
Fluorescent tubes



These tubes contain mercury vapor. Under a high voltage is applied. This excites the gas atoms so that some of the mercury atoms emit photons when they relax. When the gas relaxes it emits light. The light is higher energy than the electron beam, but the photons in the tube are of a longer wavelength. The light is emitted in all directions.

Energy levels

Electrons can be excited to higher energy levels. Either by absorbing photons from the light or by the collision. In a gas discharge tube, when an electron collides with a gas atom, it can transfer its energy to the atom.



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Xiaolong Qi



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