

Chapter 17: Chemical Equilibrium

Worksheet 17.1: K<sub>c</sub> & K<sub>p</sub>

1. For the reaction  $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$

Write the

- a)  $K_c$  expression for the reaction with  $[\text{SO}_2}]$  and  $[\text{O}_2}]$  and  $[\text{SO}_3}]$  in mol dm<sup>-3</sup>
- b)  $K_c$  expression for the reaction with  $[\text{SO}_2}]$  and  $[\text{SO}_3}]$  in mol dm<sup>-3</sup> and  $p_{\text{O}_2}$  in atm
- c)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $[\text{O}_2}]$  in mol dm<sup>-3</sup>
- d)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $p_{\text{O}_2}$  in atm
- e)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $p_{\text{O}_2}$  in atm
- f)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $p_{\text{O}_2}$  in atm
- g)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $p_{\text{O}_2}$  in atm
- h)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $p_{\text{O}_2}$  in atm
- i)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $p_{\text{O}_2}$  in atm
- j)  $K_c$  expression for the reaction with  $p_{\text{SO}_2}$  and  $p_{\text{SO}_3}$  in atm and  $p_{\text{O}_2}$  in atm

Answers

- 1a)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1b)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1c)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1d)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1e)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1f)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1g)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1h)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1i)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$
- 1j)  $K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$

# Elements Of Language First Course Answer Key

**Siddappa N.Byrareddy**



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