

# Stoichiometry of a Precipitation Reaction

Hands-On Labs, Inc.

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## Lab Report Assistant

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This document is not meant to be a substitute for a formal laboratory report. The Lab Report Assistant is simply a summary of the experiment's questions, diagrams if needed, and data tables that should be addressed in a formal lab report. The intent is to facilitate students' writing of lab reports by providing this information in an editable file which can be sent to an instructor.

## Exercise 1: Stoichiometry and a Precipitation Reaction

**Data Table 1.** Stoichiometry Values.

Initial: $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ (g)	1.00g
Initial: $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ (moles)	.0068M
Initial: $\text{CaCl}_2$ (moles)	.0068M
Initial: $\text{Na}_2\text{CO}_3$ (moles)	.0068M
Initial: $\text{Na}_2\text{CO}_3$ (g)	.8g
Theoretical: $\text{CaCO}_3$ (g)	.68g
Mass of Filter paper (g)	.9g
Mass of Filter Paper + $\text{CaCO}_3$ (g)	1.54g
Actual: $\text{CaCO}_3$ (g)	.62g
% Yield:	94.12%

## Questions

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- A. A perfect percent yield would be 100%. Based on your results, describe your degree of accuracy and suggest possible sources of error.

My accuracy was slightly off and this due to a few reasons. My scale is not the best the numbers kept switching and then it would shut off. However, I am glad that it was not too off.

Additionally, I am human and struggled with getting the perfect 1g of  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$  and I may had dropped a little without even knowing of  $\text{Na}_2\text{CO}_3$  when pouring the contents of the bag on the boat to scale. Meaning that my amount could have been off as well. But with human errors and technical errors those two could have been factors. Additionally, precipitation, my math skills, the filtering, the environment, and correct measurement could have been factors as well for causes of error in accuracy.

# Solution Stoichiometry Lab Report

**M Tight**



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