Unit Conversions

- Temperature:
$$T(°C) = T(K) - 273.15$$
$$T(°F) = 1.8 \cdot T(°C) + 32$$
$$T(°R) = T(°F) + 459.67$$
- Pressure:
$$1 \text{ bar} = 10^5 \text{ Pa} = 100 \text{ kPa}$$
$$P_{atm} = 101 \text{ kPa} = 14.7 \text{ lbf/in}^2$$
$$6895 \text{ Pa} = 1 \text{ psi}$$
- Energy + Power:
$$1 \text{ Btu} = 778 \text{ ft} \cdot \text{lbf}$$
$$1 \text{ hp} = 550 \text{ ft} \cdot \text{lbf/s} = 2545 \text{ Btu/hr}$$
$$N = \text{kg} \cdot \text{m/s}^2$$
- Volume:
$$1 \text{ m} = 3.28084 \text{ ft}$$
- Fluids:
$$\rho_{H_2O} = 1000 \text{ kg/m}^3 = 1.94 \text{ slug/ft}^3$$
$$\gamma_{H_2O} = 9800 \text{ N/m}^3 = 62.4 \text{ lbf/ft}^3$$
$$\gamma_{Hg} = 133,200 \text{ N/m}^3$$
- Mass:
$$\text{Slug} = 32.174 \text{ lb}_{mass} = \text{lb}_{mass} \cdot \text{s}^2/\text{ft}$$

Adiabatic: $\Delta Q = \Delta s = 0$ Isentropic: $\Delta s = 0$

1st Law
$$\Delta E = \Delta KE + \Delta PE + \Delta U = {}_1Q_2 - {}_1W_2$$
Rate Form: $\dot{E} = \dot{Q} - \dot{W}$

Mixtures
$$X = \frac{M_{vap}}{M} = \frac{x - x_f}{x_g - x_f} \quad (x = v, u, h)$$
$$\beta = (1 - X)\beta_f + X \cdot \beta_g$$

Ideal Gas
$$P\bar{v} = \bar{R}T, \; \bar{R} \text{ is universal}$$
$$Pv = RT, \text{ where } R = \frac{\bar{R}}{M_w}$$
$$P = \rho RT$$
$$PV = mRT$$
$$\bar{R} = \begin{cases} 8.314 \frac{kJ}{Kmol \cdot K} \\ 1.986 \frac{Btu}{lbmol \cdot °R} \\ 1545 \frac{ft \cdot lbf}{lbmol \cdot °R} \end{cases}$$

Compressed Liquid Approx.
$$x(T, P) \approx x_f(T), \; x = (v, u, h, s)$$

Polytropic Ideal Gas
$$\frac{T_2}{T_1} = \left(\frac{P_2}{P_1}\right)^{\frac{n-1}{n}} = \left(\frac{V_1}{V_2}\right)^{n-1}$$
$$_1W_2 = \begin{cases} \frac{mR(T_2 - T_1)}{(1-n)}, & n \neq 1 \\ mRT \ln\left(\frac{v_2}{v_1}\right), & n = 1 \end{cases}$$

Polytropic Process
$$P \cdot V^n = c \quad \text{or} \quad P \cdot v^n = c$$
- Volume expansion work
$$_1W_2 = \begin{cases} \frac{(P_2 V_2 - P_1 V_1)}{(1-n)}, & n \neq 1 \\ P_1 V_1 \ln\left(\frac{v_2}{v_1}\right), & n = 1 \end{cases}$$

Real gas
$$z = \frac{Pv}{RT}, \text{ where } z = \text{ compressibility factor}$$

Specific Heat (heat capacity)
C used when incompressible, otherwise C_v and C_p
$$\Delta u = \int_{T_1}^{T_2} C_v(T) \, dT$$
$$\Delta h = \int_{T_1}^{T_2} C_p(T) \, dT + v(P_2 - P_1)$$
Constant C_v approximation
$$C_v = \frac{C_v(T_1) + C_v(T_2)}{2}$$
$$C_p = C_v + R$$

Enthalpy
$$H = U + P \cdot V$$
$$h = u + P \cdot v$$

Energy Conservation
$$\frac{dE_{cv}}{dt} = \dot{Q}_{cv} - \dot{W}_{cv} + \sum_i \dot{m}_i \left(h_i + \frac{1}{2}|V_i|^2 + g \cdot z_i\right) - \sum_e \dot{m}_e \left(h_e + \frac{1}{2}|V_e|^2 + g \cdot z_e\right)$$
1 Inlet, 1 Exit: $0 = \dot{Q}_{cv} - \dot{W}_{cv} + \dot{m}\left[(h_i - h_e) + \frac{1}{2}(|V_i|^2 - |V_e|^2) + g(z_i - z_e)\right]$

Mass Conservation
$$\frac{dMass_{cv}}{dt} = \sum \dot{m}_i - \sum \dot{m}_e$$
$$\dot{m} = \rho A |V| = \frac{A|V|}{v}$$
Volumetric flow rate $= A|V|$
Flow work $= (P \cdot A) \cdot |V|$

Entropy
$$dS = \left(\frac{\delta Q}{T}\right)_{rev}, \; S = s \cdot m$$
$$S_2 - S_1 = \int_1^2 \left(\frac{\delta Q}{T}\right)_b$$
$$T \cdot ds = du + P \cdot dv$$
$$T \cdot ds = dh - v \cdot dP$$
Rev: $_1Q_2 = \int_1^2 T ds = mT(s_2 - s_1)$

Entropy Balance (Cycle)
$$\frac{dS_{cv}}{dt} = \sum \frac{\dot{Q}}{T} + \sum \dot{m}_i \cdot s_i - \sum \dot{m}_e \cdot s_e + \dot{\sigma}_{cv}$$
1 in, 1 out: $0 = \sum \frac{\dot{Q}}{T} + \dot{m}(s_1 - s_2) + \dot{\sigma}_{cv}$
$$\Delta s = s_2 - s_1 = \frac{1}{\dot{m}}\left(\sum \frac{\dot{Q}}{T} + \dot{\sigma}_{cv}\right)$$

Entropy and Ideal Gas
$$\Delta s = s(T_2, P_2) = s(T_1, P_1) = \int_{T_1}^{T_2} \frac{C_p(T)}{T} dT + R \cdot \ln\left(\frac{v_2}{v_1}\right)$$
$$\Delta s = s(T_2, P_2) - s(T_1, P_1) = \int_{T_1}^{T_2} \frac{C_p(T)}{T} dT - R \cdot \ln\left(\frac{P_2}{P_1}\right)$$
$$\Delta s = s°(T_2) - s°(T_1) - R \cdot \ln\left(\frac{P_2}{P_1}\right)$$

2nd Law and Cycle COP's
$$COP_{irrev} < COP_{rev} = COP_{max}$$
$$COP_{rev,1} = COP_{rev,2}$$
Power: $\eta_{rev} = \frac{W}{Q_H} = \frac{Q_H - Q_L}{Q_H} = 1 - \frac{T_L}{T_H}$
Refrig: $\beta_{rev} = \frac{Q_{in}}{W} = \frac{Q_{in}}{Q_{out} - Q_{in}} = \frac{T_C}{T_H - T_C}$
Heat Pump: $\gamma_{rev} = \frac{Q_{out}}{W} = \frac{Q_{out}}{Q_{out} - Q_{in}} = \frac{T_H}{T_H - T_C}$

2nd Law and Entropy
$$\oint_{cycle} \left(\frac{\delta Q}{T}\right)_b \leq 0$$
S production: $\sigma = -\oint \left(\frac{\delta Q}{T}\right)_b \geq 0$
$\sigma > 0$: irreversible cycle
$\sigma = 0$: reversible cycle
$\sigma < 0$: impossible

Work and Power
$$W = \int F \cdot dr$$
- Volume expansion work:
$$W = \int P \, dV$$
$$W = m \cdot \int P \, dv$$
$$Power = \frac{dW}{dt}$$

Isentropic Efficiency
Turbine: $\eta_t = \frac{\dot{W}_t/\dot{m}}{(\dot{W}_t/\dot{m})_s} = \frac{h_1 - h_2}{h_1 - h_{2s}} \leq 1$
Nozzle: $\eta_n = \frac{\frac{1}{2}|V_2|^2}{(\frac{1}{2}|V_2|^2)_s} \leq 1$
Compressor/Pump: $\eta_{c/p} = \frac{(\dot{W}_c/\dot{m})_s}{(\dot{W}_c/\dot{m})} \leq 1$

Entropy Balance (Process)
$$S_2 - S_1 = \int_1^2 \left(\frac{\delta Q}{T}\right)_b + \sigma$$
$$S_2 - S_1 = \sum \frac{Q}{T} + \sigma$$
$$\frac{dS}{dt} = \sum \frac{\dot{Q}}{T} + \dot{\sigma}$$

Energy
$$KE = \frac{1}{2}mv^2 \qquad PE = mgz$$

1st Law and Cycles
Power: $Q_H - Q_C = W_{cycle}$
R+HP: $Q_{out} - Q_{in} = W_{cycle}$

Velocity^2 to kJ/kg
$$\frac{m^2}{s^2} \cdot \frac{N \cdot s^2}{kg \cdot m} \cdot \frac{J}{N \cdot m} \cdot \frac{kJ}{10^3 J}$$

Engineering Thermodynamics Equation Sheet

Frank A Di Bella

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