

9-2 Review and Reinforcement

Chemical Equations

Refer to the unbalanced chemical equation below to answer each of the following questions. Write your answers in the space provided.



- Write the names of the reactants in this equation.
Aluminum Nitrate + Iron (II) chloride
- Write the names of the products in this equation.
Iron (II) Nitrate + Aluminum chloride
- What would the coefficient of $\text{Al(NO}_3)_3$ be in the balanced equation? Two
- What would the coefficient of $\text{Fe(NO}_3)_2$ be in the balanced equation? Three
- What would the coefficient of AlCl_3 be in the balanced equation? Two
- How many oxygen atoms would there be on each side of the balanced equation?
18

Answer each of the following questions in the space provided.

- What should you do if an equation you have written cannot be balanced?
make sure reactants + products are written correctly, try again, ask for help
- How is balancing a chemical reaction related to the law of conservation of matter?
mass of reactants must = mass of products
- Describe the correct way to balance an equation.
add only coefficients never change formulas or subscripts
- How does a chemical equation relate the macroscopic, observable reaction to the interactions occurring at the microscopic level?
Same

3 1 Review Reinforcement Answer Key

**Raymond Barnett, Michael Ziegler, Karl
Byleen, David Sobecki**

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