

## Lab 22A

### Investigating Chemical Equilibrium

**Purpose:** To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and make use of Le Chatelier's analysis and an equilibrium shift construction through applying Le Chatelier's principle.

#### Proble:

1. Write a balanced equation for the reaction in part 1.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.  
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{FeSCNH}^{3+}(\text{aq}) + \text{HSCN}(\text{aq})$$
4. In part 4, what is the conjugate acid of the base  $\text{FeSCN}^{2+}$  and what is the  $\text{pH}$  of the solution? Explain why there are the color change and how the color change is related to the reaction. The  $\text{pH}$  of the solution during the reaction is constant. The color of the solution and the color of the solution are constant. Explain why there are the color change.
5. What is the conjugate acid of the base  $\text{FeSCN}^{2+}$ ?  
There is a color change in the solution. The color of the solution is constant. Explain why there are the color change.
6. What is the conjugate acid of the base  $\text{FeSCN}^{2+}$ ?  
There is a color change in the solution. The color of the solution is constant. Explain why there are the color change.
7. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
$$\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$$
8. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
$$\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$$
9. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
$$\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$$
10. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
$$\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$$

# Investigating Chemical Equilibrium Lab 12a Answer

**Emilie Sanchez**



## Investigating Chemical Equilibrium Lab 12a Answer:

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