

Fuel Cells Harvest Chemical Energy

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The fuel cell converts chemical energy from a hydrocarbon fuel (such as hydrogen, butane, and propane) into electrical energy. The fuel cell is a device that converts chemical energy from a hydrocarbon fuel (such as hydrogen, butane, and propane) into electrical energy. The fuel cell is a device that converts chemical energy from a hydrocarbon fuel (such as hydrogen, butane, and propane) into electrical energy. The fuel cell is a device that converts chemical energy from a hydrocarbon fuel (such as hydrogen, butane, and propane) into electrical energy.

Applying Your Knowledge

Answer the questions below and complete additional responses to Student Activities below.

1. Write a balanced chemical equation for the combustion of propane. $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

2. Write a balanced chemical equation for the combustion of butane. $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$

3. Write a balanced chemical equation for the combustion of propane. $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

4. Write a balanced chemical equation for the combustion of butane. $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$

5. Write a balanced chemical equation for the combustion of propane. $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

6. Write a balanced chemical equation for the combustion of butane. $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$

7. Write a balanced chemical equation for the combustion of propane. $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

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9. Write a balanced chemical equation for the combustion of propane. $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

10. Write a balanced chemical equation for the combustion of butane. $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$

11. Write a balanced chemical equation for the combustion of propane. $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

12. Write a balanced chemical equation for the combustion of butane. $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$

13. Write a balanced chemical equation for the combustion of propane. $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

14. Write a balanced chemical equation for the combustion of butane. $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$

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Kathleen Armour



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