

9

CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253–258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

► Monatomic Ions (pages 253–256)

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

- a. alkali metal
- b. alkaline earth metal
- c. transition metal
- d. nonmetal

5. The _____ *Stock system* _____ of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix *-ous* to name the cation with the _____ *lesser* _____ charge, and the suffix *-ic* to name the cation with the _____ *greater* _____ charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

Chapter 9 Chemical Names And Formulas Section Review Answers

Shasha Hu



Chapter 9 Chemical Names And Formulas Section Review Answers:

Essential Concepts of Chemistry Sharon Sherman, Alan Sherman, 1999 Designed especially for students who have little or no background in chemistry or mathematics Essential Concepts of Chemistry makes complex concepts understandable This text provides an inexpensive one color alternative for introductory chemistry courses and emphasizes everyday applications of chemistry

Understanding Chemistry Fred M. Dewey, 1994 *Chemistry* Henry Dorin, 1987

Chemistry George Tyler Miller, 1987 **Essentials of Chemistry** Dennis D. Staley, Antony C. Wilbraham, 1984

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Prentice Hall Chemistry**, 2000

Fundamentals of Chemistry Ralph A. Burns, 2003 For one semester preparatory chemistry courses or general purpose introductory chemistry courses This clearly written well illustrated versatile textbook provides thorough coverage of chemistry with a balance of problem solving skills real world applications and an emphasis on critical thinking and the process of science A supporting theme throughout the text continually emphasizes that chemistry is everywhere

Holt Chemistry Ralph Thomas Myers, 2004 **Foundations of College Chemistry, International Adaptation** Morris Hein, Susan Arena, Cary Willard, 2023-05-16 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity

Introductory Chemistry Victor S. Krimsley, Darold E. Skerritt, Beverly B. Harrison, 1986 This pedagogically rich text has all the necessary features to hook introductory students and keep them interested and successful in preparatory chemistry This book carefully guides beginning students through the fundamental principals and calculations required for their subsequent success in either the general chemistry course or the short organic and biochemistry course Krimsley while dedicated to conceptual understanding and skill building presents a solid book that provides students with complete explanations on every point and helps them work through the material methodically with many examples and hints His approach is cohesive and coherent always patient and interesting Krimsley introduces all topics through an example students are already familiar with He continually reminds them of objectives and provides many opportunities to practice and then check their answers The text begins with a study of atomic and molecular structure prior

to treating the various categories of chemical reactions The organization is designed to get students into chemistry quickly yet methodically The classification of matter begins in Chapter 2 before the chapter on measurement and the coverage of bonding appears in Chapter 8 The elementary concepts of chemistry are presented with an emphasis on mathematical calculations and problem solving strategies *Chemistry* James E. Brady, Fred Senese, 2004-02-04 Publisher Description

Foundations of College Chemistry Morris Hein, 1991 Features strong problem solving and carefully constructed problem sets **Basic Chemistry** Steven S. Zumdahl, 1999 For a full description see catalog entry for Zumdahl Introductory Chemistry A Foundation 4 e **Marketing Information Guide** ,1964 Machinery ,1918 **Railroad Gazette** ,1908

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