

10

CHEMICAL QUANTITIES

Chapter Test A

A. Matching

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

Column A

- _____ 1. standard temperature and pressure
- _____ 2. molar volume
- _____ 3. representative particle
- _____ 4. molar mass
- _____ 5. atomic mass
- _____ 6. Avogadro's number
- _____ 7. mole (mol)
- _____ 8. empirical formula
- _____ 9. percent composition
- _____ 10. molar mass

Column B

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i. 6.02×10^{23} representative particles of a substance
- j. 0°C and 101.3 kPa

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol, $\text{C}_3\text{H}_7\text{O}$?
- a. $4 \times (6.02 \times 10^{23})$
- b. 4
- c. 32
- d. $32 \times (6.02 \times 10^{23})$
- _____ 12. The mass of a mole of NaCl is its:
- a. molar mass.
- b. empirical formula.
- c. gram atomic mass.
- d. atomic mass.

Chemical Quantities Chapter 10 Test Answer Key

Kathleen Armour



Chemical Quantities Chapter 10 Test Answer Key:

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General, Organic, & Biological Chemistry Janice Gorzynski Smith,2024-12 My goal in writing this text was to relate the fundamental concepts of general organic and biological chemistry to the world around us and in this way illustrate how chemistry explains many aspects of everyday life A key feature is the use of molecular art to illustrate and explain common phenomena we encounter every day Each topic is broken down into small chunks of information that are more manageable and easily learned Students are given enough detail to understand basic concepts such as how soap cleans away dirt and why trans fats are undesirable in the diet without being overwhelmed **Lessons in Qualitative and Volumetric Chemical Analysis** Charles O. Curtman,1894 Code of Federal Regulations, Title 40, Protection of Environment, Pt. 790-End, Revised as of July 1 2005 ,2005-10 The Code of Federal Regulations is a codification of the general and permanent rules published in the Federal Register by the Executive departments and agencies of the United States Federal Government

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