

Chapter Test B

Chapter: Chemical Equations and Reactions

Match a name space (possibly writing the name of the react or product molecule) completed with reactants/products atoms/molecules.

- The product(s) of a chemical reaction will be represented by a double arrow pointing towards the right. The substance(s) on the left side of the reaction is/are
 - reactant(s)**
 - unbalanced equation
 - product
- An balanced chemical equation is represented by adding the
 - coefficients**
 - subscripts
 - structure of the molecules
 - number of products
- An unbalanced equation that is written as $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ is
 - correct
 - incorrect because**
 - unbalanced
- The balancing for unbalanced equation on atoms is done as follows
 - identify reactants and products for the unbalanced equation
 - count atoms for each
 - write unbalanced equation
 - compare reactants and products for the unbalanced equation**
- The coefficients for balanced equation
 - indicate the number of molecules of each substance that react**
 - show the number of grams of each substance that react
 - show the number moles of the substances
 - show the relative proportions for each atom
- If $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ then react H_2 is the same as 2H_2 , then
 - it will replace 2H_2 in the equation**
 - H_2 will replace 2H_2 in the equation
 - H_2 will replace 2H_2 in the equation
 - H_2 will replace 2H_2 in the equation

Chapter 8 Test B

Sturgiss Brown Davis



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Test Bank Checkett,2000-11 *Science in Your World: Teacher resource masters* Jay K. Hackett,1991

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