

CHEMISTRY



Properties of Matter

- Matter is anything that has mass and occupies space.
- Matter is made up of particles.
- Particles are very small.
- Particles are continuously moving.
- Particles attract each other.

States of Matter

- Solid: Particles are closely packed and have a fixed shape and volume.
- Liquid: Particles are less closely packed and have a fixed volume but no fixed shape.
- Gas: Particles are far apart and have no fixed shape or volume.

Changes of State

- Melting: Solid to liquid.
- Freezing: Liquid to solid.
- Boiling: Liquid to gas.
- Condensation: Gas to liquid.
- Sublimation: Solid to gas.
- Deposition: Gas to solid.

Atomic Structure

Atoms are the smallest particles that cannot be created or destroyed in a chemical reaction. They are made up of three sub-particles: electrons, protons, and neutrons.

Electrons

- Discovered by Cathode ray experiment.
- Charge: -1.6×10^{-19} C
- Mass: 9.1×10^{-31} kg

Protons

- Discovered by Rutherford's alpha particle scattering experiment.
- Charge: $+1.6 \times 10^{-19}$ C
- Mass: 1.67×10^{-27} kg

Neutrons

- Discovered by Chadwick.
- Charge: 0
- Mass: 1.67×10^{-27} kg

Atomic Structure

Atoms are made up of a central nucleus containing protons and neutrons, surrounded by electrons. The nucleus is positively charged, while the electrons are negatively charged. The overall atom is electrically neutral.

Atomic Number

The atomic number of an element is the number of protons present in the nucleus of an atom of that element. It is denoted by the symbol Z .

Mass Number

The mass number of an element is the sum of the number of protons and neutrons present in the nucleus of an atom of that element. It is denoted by the symbol A .

Element	Atomic Number (Z)	Mass Number (A)
Hydrogen	1	1
Helium	2	4
Lithium	3	7
Boron	5	11
Carbon	6	12
Nitrogen	7	14
Oxygen	8	16
Fluorine	9	19
Neon	10	20

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