

## Colorimetric Analysis Tutorial Answers

1. a) The limitations of the Beer-Lambert Law are:

- Deviations in absorption coefficients at high concentrations (>0.01M) due to electrostatic interactions between molecules in close proximity
- Scattering of light due to particulates in the sample
- Fluorescence or phosphorescence of the sample
- Refractive index changes with analyte concentration
- Shifts in chemical equilibrium as a function of concentration
- Non-monochromatic radiation (deviations can be minimized by using a relatively flat part of the absorption spectrum such as the maximum of an absorption band)
- Stray light
- Unmatched cells

See D A Skoog, F J Holler, S R Crouch, (2007) "Principles of Instrumental Analysis", 6th Ed., p 336-343 for a full description.

b) Colorimetric analysis of mixtures requires/assumes no interaction between the various species present i.e. they do not change each other's colour/absorbance spectrum.

2. Break the problem down into steps and use equations for propagation of errors.

	Value	SD
$A_1$	0.312	0.003
$A_2$	0.092	0.002
$l$	1.000	0.005
$\epsilon_{A1}$	70.660	2.165
$\epsilon_{A2}$	3.510	0.197
$\epsilon_{B1}$	16.910	0.786
$\epsilon_{B2}$	107.230	3.055

Step	Value	Error
$\epsilon_{B1} A_2$	1.5557	0.0798
$\epsilon_{B2} A_1$	33.4558	1.0060
$\epsilon_{A2} \epsilon_{B1}$	59.3541	4.3254
$\epsilon_{A1} \epsilon_{B2}$	7576.8718	317.0067
$\epsilon_{B1} A_2 - \epsilon_{B2} A_1$	-31.9000	1.0091
$\epsilon_{A2} \epsilon_{B1} - \epsilon_{A1} \epsilon_{B2}$	7517.5177	317.0362
$c_A$	0.004243	0.0002247

$$= 1.5557 \times \sqrt{[(0.786/16.91)^2 + (0.002/0.092)^2]}$$

$$= 33.4558 \times \sqrt{[(3.055/107.23)^2 + (0.003/0.312)^2]}$$

$$= 59.3541 \times \sqrt{[(0.197/3.51)^2 + (0.786/16.91)^2]}$$

$$= 7576.8718 \times \sqrt{[(2.165/70.66)^2 + (3.055/107.23)^2]}$$

$$= \sqrt{[(0.0798)^2 + (1.0060)^2]}$$

$$= \sqrt{[(4.3254)^2 + (317.0067)^2]}$$

$$= 0.004243 \times \sqrt{[(1.0091/-31.9000)^2 + (317.0362/-7517.5177)^2 + (0.005/1)^2]}$$

Step	Value	Error
$\epsilon_{A1} A_2$	1.0951	0.0624
$\epsilon_{A2} A_1$	6.5007	0.2442
$\epsilon_{B1} \epsilon_{A2}$	59.3541	4.3254
$\epsilon_{B2} \epsilon_{A1}$	7576.8718	317.0067
$\epsilon_{A1} A_2 - \epsilon_{A2} A_1$	-5.4056	0.2521
$\epsilon_{B2} \epsilon_{A1} - \epsilon_{B1} \epsilon_{A2}$	-7517.5177	317.0362
$c_B$	0.000719	0.000045

Therefore:  $c_A = 0.00424 \pm 0.00022$  M and  $c_B = 0.00072 \pm 0.00005$  M

# Colorimetric Analysis Lab Answers

**Robert R. Redfield**



## **Colorimetric Analysis Lab Answers:**

*Practical Instrumental Analysis* Sergio Petrozzi, 2012-09-24 This practical book in instrumental analytics conveys an overview of important methods of analysis and enables the reader to realistically learn the principally technology independent working techniques the analytical chemist uses to develop methods and conduct validation What is to be conveyed to the student is the fact that analysts in their capacity as problem solvers perform services for certain groups of customers i e the solution to the problem should in any case be processed in such a way as to be fit for purpose The book presents sixteen experiments in analytical chemistry laboratory courses They consist of the classical curriculum used at universities and universities of applied sciences with chromatographic procedures atom spectrometric methods sensors and special methods e g field flow fractionation flow injection analysis and N determination according to Kjeldahl The carefully chosen combination of theoretical description of the methods of analysis and the detailed instructions given are what characterizes this book The instructions to the experiments are so detailed that the measurements can for the most part be taken without the help of additional literature The book is complemented with tips for effective literature and database research on the topics of organization and the practical workflow of experiments in analytical laboratory on the topic of the use of laboratory logs as well as on writing technical reports and grading them Evaluation Guidelines for Laboratory Experiments A small introduction to Quality Management a brief glance at the history of analytical chemistry as well as a detailed appendix on the topic of safety in analytical laboratories and a short introduction to the new system of grading and marking chemicals using the Globally Harmonized System of Classification and Labelling of Chemicals GHS round off this book This book is therefore an indispensable workbook for students internship assistants and lecturers in the area of chemistry biotechnology food technology and environmental technology in the basic training program of analytics at universities and universities of applied sciences

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**Photometric Chemical Analysis** John Howe Yoe, 1928

[Colorimetric Analysis](#) Foster Dee Snell, 1921

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## **Table of Contents Colorimetric Analysis Lab Answers**

1. Understanding the eBook Colorimetric Analysis Lab Answers
  - The Rise of Digital Reading Colorimetric Analysis Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Colorimetric Analysis Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Colorimetric Analysis Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Colorimetric Analysis Lab Answers
  - Personalized Recommendations
  - Colorimetric Analysis Lab Answers User Reviews and Ratings
  - Colorimetric Analysis Lab Answers and Bestseller Lists
5. Accessing Colorimetric Analysis Lab Answers Free and Paid eBooks
  - Colorimetric Analysis Lab Answers Public Domain eBooks
  - Colorimetric Analysis Lab Answers eBook Subscription Services

- Colorimetric Analysis Lab Answers Budget-Friendly Options
- 6. Navigating Colorimetric Analysis Lab Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Colorimetric Analysis Lab Answers Compatibility with Devices
  - Colorimetric Analysis Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Colorimetric Analysis Lab Answers
  - Highlighting and Note-Taking Colorimetric Analysis Lab Answers
  - Interactive Elements Colorimetric Analysis Lab Answers
- 8. Staying Engaged with Colorimetric Analysis Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Colorimetric Analysis Lab Answers
- 9. Balancing eBooks and Physical Books Colorimetric Analysis Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Colorimetric Analysis Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Colorimetric Analysis Lab Answers
  - Setting Reading Goals Colorimetric Analysis Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Colorimetric Analysis Lab Answers
  - Fact-Checking eBook Content of Colorimetric Analysis Lab Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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