

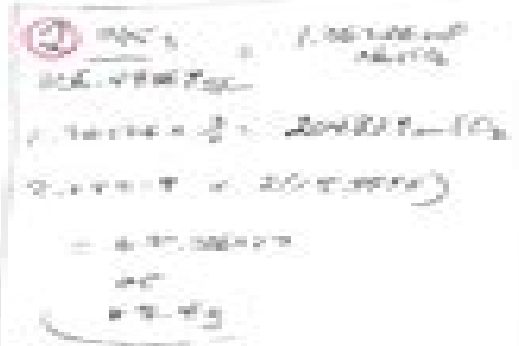
KEY

1. Complete and classify the following partial equations

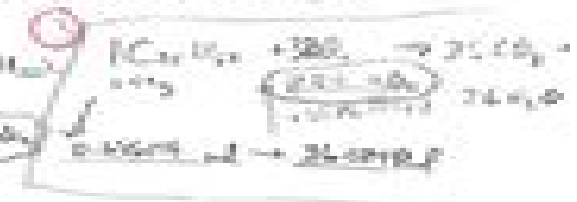
- a) $1Fe + 3H_2SO_4 \rightarrow 1FeSO_4 + 3H_2$ (single repl)
- b) $2Na + Cl_2 \rightarrow 2NaCl$
- c) $1Zn + 1Pb(NO_3)_2 \rightarrow 1Zn(NO_3)_2 + 1Pb$ (single repl)
- d) $Al(OH)_3 + HCl \rightarrow AlCl_3 + 3H_2O$ (acid base)
- e) $2Cu + O_2 \rightarrow 2CuO$ (synthesis)
- f) $4Cu + 1O_2 \rightarrow 2Cu_2O$ (synthesis)
- g) $Al_2O_3 \rightarrow 2Al + 3/2O_2$ (decomposition)
- h) $1N_2 + 3H_2 \rightarrow 2NH_3$ (synthesis)
- i) $C_2H_6 + O_2 \rightarrow 2CO_2 + 3H_2O$ (combustion)
- j) $1Mg + 2HNO_3 \rightarrow 1Mg(NO_3)_2 + H_2$ (single repl)
- k) $Na + H_2O \rightarrow NaOH + H_2$ (single repl)
- l) $3NaOH + H_3PO_4 \rightarrow 1Na_3PO_4 + 3H_2O$ (acid base)
- m) $H_2SO_4 + Al(OH)_3 \rightarrow Al_2(SO_4)_3 + 6H_2O$ (acid base)
- n) $2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$ (combustion)
- o) $AlCl_3 \rightarrow Al + 3/2Cl_2$ (decomposition)
- p) $2KOH + H_2SO_4 \rightarrow 1K_2SO_4 + 2H_2O$ (acid base)
- q) $C_2H_6 + O_2 \rightarrow 2CO_2 + 3H_2O$ (combustion)
- r) $1NaOH + HCl \rightarrow 1NaCl + 1H_2O$ (acid base)
- s) $NaHCO_3 + H_2SO_4 \rightarrow Na_2CO_3 + H_2O + CO_2$ (acid base)
- t) $1CH_4 + 2O_2 \rightarrow 1CO_2 + 2H_2O$ (combustion)
- u) $H_2O \rightarrow H_2 + 1/2O_2$ (decomposition)
- v) $2HNO_3 + CuCO_3 \rightarrow 1Cu(NO_3)_2 + 1H_2O + CO_2$ (acid base)
- w) $CH_4 + O_2 \rightarrow CO_2 + 2H_2O$ (combustion)
- x) $Pb + H_2SO_4 \rightarrow 1PbSO_4 + H_2$ (single repl)
- y) $KI + Br_2 \rightarrow 2KBr + I_2$ (single repl)
- z) $1Ba(OH)_2 + 2HNO_3 \rightarrow 1Ba(NO_3)_2 + 2H_2O$ (acid base)
- aa) $Pb + Cu(NO_3)_2 \rightarrow Pb(NO_3)_2 + Cu$ (single repl)



1
2
3
4
5



- 2. Given: $2NaClO_3 \rightarrow 2NaCl + 3O_2$. Calculate the mass of oxygen gas produced if 143 g of $NaClO_3$ is decomposed. **65.4g**
- 3. Calculate the mass of oxygen gas required to burn 242 g of wax ($C_{21}H_{42}$). **254g**
- 4. If 1.40 moles of butane (C_4H_{10}) is burned:
 - a) calculate the number of moles of oxygen gas required. **9.80 mol**
 - b) calculate the number of moles of carbon dioxide produced. **5.60 mol**
 - c) calculate the mass of water produced. **10.8 mol**
- 5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.
 - a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. **27.4g Al**
 - b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. **192g Zn**
- 6. Given: $Cu + 2AgNO_3 \rightarrow Cu(NO_3)_2 + 2Ag$
 - a) How many grams of Cu are required to react with 3.00 g of silver nitrate? **0.544g Cu**
 - b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? **1.70g Ag**



Nelson Chemistry 30 Answer Key

Lingjun Ying

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Table of Contents Nelson Chemistry 30 Answer Key

1. Understanding the eBook Nelson Chemistry 30 Answer Key
 - The Rise of Digital Reading Nelson Chemistry 30 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 30 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 30 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 30 Answer Key
 - Personalized Recommendations
 - Nelson Chemistry 30 Answer Key User Reviews and Ratings
 - Nelson Chemistry 30 Answer Key and Bestseller Lists
5. Accessing Nelson Chemistry 30 Answer Key Free and Paid eBooks
 - Nelson Chemistry 30 Answer Key Public Domain eBooks
 - Nelson Chemistry 30 Answer Key eBook Subscription Services
 - Nelson Chemistry 30 Answer Key Budget-Friendly Options
6. Navigating Nelson Chemistry 30 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Nelson Chemistry 30 Answer Key Compatibility with Devices
 - Nelson Chemistry 30 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nelson Chemistry 30 Answer Key
 - Highlighting and Note-Taking Nelson Chemistry 30 Answer Key
 - Interactive Elements Nelson Chemistry 30 Answer Key

8. Staying Engaged with Nelson Chemistry 30 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 30 Answer Key
9. Balancing eBooks and Physical Books Nelson Chemistry 30 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 30 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nelson Chemistry 30 Answer Key
 - Setting Reading Goals Nelson Chemistry 30 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nelson Chemistry 30 Answer Key
 - Fact-Checking eBook Content of Nelson Chemistry 30 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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