

KEY

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of K_a for the acid HBrO. (4 marks) $\text{pH} = 4.65$ so $[\text{H}_3\text{O}^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} \text{ M}$

$$\text{HBrO} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{BrO}^-$$

[I]	0.25		0	0
[C]	-2.239×10^{-5}		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		2.239×10^{-5}	2.239×10^{-5}

④

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{BrO}^-]}{[\text{HBrO}]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) $\text{pH} = 2.355$ $[\text{H}_3\text{O}^+] = \text{antilog}(-2.355) = 0.004416 \text{ M}$

$$\text{C}_6\text{H}_5\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{C}_6\text{H}_5\text{COO}^-$$

[I]	C_0		0	0
[C]	-0.004416		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		0.004416	0.004416

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_6\text{H}_5\text{COO}^-]}{[\text{C}_6\text{H}_5\text{COOH}]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 \text{ M}$$

Answer $C_0 = 0.30 \text{ M}$

5. Find the value of K_b for the oxalate ion ($\text{C}_2\text{O}_4^{2-}$). (1 mark)

$$K_b(\text{C}_2\text{O}_4^{2-}) = \frac{K_w}{K_a(\text{HC}_2\text{O}_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

①

Answer $K_b = 1.6 \times 10^{-10}$

6. The value of K_b for the weak base methylamine (CH_3NH_2) is 4.4×10^{-4} . Calculate the value of K_a for the acid CH_3NH_3^+ . (1 mark)

$$K_a(\text{CH}_3\text{NH}_3^+) = \frac{K_w}{K_b(\text{CH}_3\text{NH}_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

①

Answer $K_a = 2.3 \times 10^{-11}$

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Table of Contents Basic Stoichiometry Phet Lab Answer Key

1. Understanding the eBook Basic Stoichiometry Phet Lab Answer Key
 - The Rise of Digital Reading Basic Stoichiometry Phet Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Basic Stoichiometry Phet Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Basic Stoichiometry Phet Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Stoichiometry Phet Lab Answer Key
 - Personalized Recommendations

- Basic Stoichiometry Phet Lab Answer Key User Reviews and Ratings
- Basic Stoichiometry Phet Lab Answer Key and Bestseller Lists
- 5. Accessing Basic Stoichiometry Phet Lab Answer Key Free and Paid eBooks
 - Basic Stoichiometry Phet Lab Answer Key Public Domain eBooks
 - Basic Stoichiometry Phet Lab Answer Key eBook Subscription Services
 - Basic Stoichiometry Phet Lab Answer Key Budget-Friendly Options
- 6. Navigating Basic Stoichiometry Phet Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Basic Stoichiometry Phet Lab Answer Key Compatibility with Devices
 - Basic Stoichiometry Phet Lab Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Basic Stoichiometry Phet Lab Answer Key
 - Highlighting and Note-Taking Basic Stoichiometry Phet Lab Answer Key
 - Interactive Elements Basic Stoichiometry Phet Lab Answer Key
- 8. Staying Engaged with Basic Stoichiometry Phet Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Basic Stoichiometry Phet Lab Answer Key
- 9. Balancing eBooks and Physical Books Basic Stoichiometry Phet Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Basic Stoichiometry Phet Lab Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Basic Stoichiometry Phet Lab Answer Key
 - Setting Reading Goals Basic Stoichiometry Phet Lab Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Basic Stoichiometry Phet Lab Answer Key
 - Fact-Checking eBook Content of Basic Stoichiometry Phet Lab Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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