

Chapter 4 – Atomic Structure

Short Answer: Answer the following questions in complete sentences using science vocabulary. Make sure that your answers are complete and detailed.

1. A sample of calcium contains calcium-40, calcium-44, calcium-42, calcium-48, calcium-43, and calcium-46 atoms. Explain why these atoms have different mass numbers but must have the same atomic number.

The atoms have different mass numbers because they have a different number of neutrons. The mass number is equaled to the number of protons plus neutrons.

The atoms must have the same atomic number because the atomic number is equaled to the number of protons. The number of protons is unique for each element.

2. Complete the following table. Circle the name of the element(s) that are isotopes. Complete Bohr's Model diagrams for the element with eleven electrons and the element with an atomic number of 9.

Element Name	Element Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Scandium	Sc	21	45	21	24	21
Fluorine	F	9	19	9	10	9
Sodium	Na	11	24	11	13	11
Answers Vary					82	

Chapter 4 Atomic Structure Worksheet Answers

RJ Shavelson



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interesting comprehensive and appealing Diagrams illustrations and text have been integrated to enhance comprehension Definitions and other important scientific information are highlighted Throughout the series investigations related to the text enable the learners to learn through experimentation Quick revision of each chapter has been given under the caption Highlights in Review Online Support It provides Video lectures Unit wise interactive exercises Chapterwise Worksheet Solution of textbook questions for Teachers only E Book for Teachers only I hope this series would meet the needs and requirements of the curriculum to achieve the learning outcomes as laid down in the curriculum Suggestions and constructive feedback for the further improvement of the book shall be gratefully acknowledged and incorporated in the future edition of the book Author *Earth Science* Samuel N. Namowitz, Nancy E. Spaulding, 1989 **Science of Life: Biology Parent Lesson Plan**, 2013-08-01 The Science of Life Biology Course Description This is the suggested course sequence that allows one core area of science to be studied per semester You can change the sequence of the semesters per the needs or interests of your student materials for each semester are independent of one another to allow flexibility Semester 1 Intro to Science Have you ever wondered about human fossils cave men skin color ape men or why missing links are still missing Want to discover when T Rex was small enough to fit in your hand Or how old dinosaur fossils are and how we know the age of these bones Learn how the Bibles world view not evolution s unites evidence from science and history into a solid creation foundation for understanding the origin history and destiny of life including yours In Building Blocks in Science Gary Parker explores some of the most interesting areas of science fossils the errors of evolution the evidences for creation all about early man and human origins dinosaurs and even races Learn how scientists use evidence in the present how historians use evidence of the past and discover the biblical world view not evolution that puts the two together in a credible and scientifically sound way Semester 2 Life Science Study clear biological answers for how science and Scripture fit together to honor the Creator Have you ever wondered about such captivating topics as genetics the roll of natural selection embryonic development or DNA and the magnificent origins of life Within Building Blocks in Life Science you will discover exceptional insights and clarity to patterns of order in living things including the promise of healing and new birth in Christ Study numerous ways to refute the evolutionary worldview that life simply evolved by chance over millions of years The evolutionary worldview can be found filtered through every topic at every age level in our society It has become the overwhelmingly accepted paradigm for the origins of life as taught in all secular institutions This dynamic education resource helps young people not only learn science from a biblical perspective but also helps them know how to defend their faith in the process **Building Block** Prentice-Hall Staff, 1994 [Exercises for Principles of Imaging Science and Protection](#) Michael A. Thompson, Steven B. Dowd, Janice D. Hall, Marian P. Hattaway, 1994 A smart collection of additional review questions and practise problems to further reinforce basic concepts The laboratory manual section includes about 30 basic experiments which concentrate more on radiographic factors which affect the quality of the resulting radiograph Readers

with a weaker math science background will appreciate the remediation sections and self tests conclude each section to gauge the reader's weaknesses *Secondary Teaching Methods* Kenneth T. Henson, 1981 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Discover Science: Teacher's annotated edition*, 1991 Science content helps develop the skills needed to understand how science works learn new concepts solve problems and make decisions in today's technological society *Special Events* Alan L. Wendroff, 1999-01-28 Providing a concrete outline on event planning within the larger framework for achieving the organizational goals this book introduces the building blocks for a successful special event or any fund raising programme including raising money marketing and public relations building and expanding the non profit donor base building and motivating the volunteer pool training and motivating your board to solicit funds and educating the public about your organization and its mission **Te HS&T J** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004-02 **Holt Science and Technology** Holt Rinehart & Winston, 2004-02 *Chemistry* William L. Masterton, 1996-08 This new edition of CHEMISTRY PRINCIPLES AND REACTIONS continues to provide students with the core material essential to understanding the principles of general chemistry Masterton and Hurley cover the basics without sacrificing the essentials appealing to several markets Appropriate for either a one or two semester course CHEMISTRY PRINCIPLES AND REACTIONS Fifth Edition is three hundred pages shorter than most general chemistry texts and lives up to its long standing reputation as THE student oriented text Though this text is shorter in length than most other General Chemistry books it is not lower in level and with the addition of the large volume of content provided by the revolutionary GENERAL CHEMISTRY INTERACTIVE 3 0 CD ROM that is included with every copy it has a depth and breadth rivaling much longer books **Te HS&T 2007 Shrt Crs M** Holt Rinehart & Winston, 2007 **Glencoe Science**, 2002

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