

## Chapter 4 Section 2 Notes

### Combination of Atoms

- **compound** – a substance made up of atoms of two or more different elements chemically combined
- **properties of a compound differ from those of the individual elements (NaCl)**

### Molecules

- **molecule** – a group of atoms that are held together by chemical forces; a molecule is the smallest unit of matter that can exist by itself and retain all of a substance's chemical properties
- **diatomic molecules** – made up of only two atoms of the same element
  - o  $H_2, O_2, F_2, I_2, Br_2, Cl_2, N_2$

### Chemical Formulas

- **elements occur in the same relative proportions**
- **chemical formula** – a combination of letters and number that show types and amounts of elements
- **water**
  - o  $H_2O$
  - o **Subscript**
    - **small number in front of an element**
    - **related to the preceding element**
    - **identifies how many of that element**
  - o **two hydrogen atoms**
  - o **one oxygen atom**

### Chemical Equations

- **a formula that describes a reaction**
  - o  $H_2 + O_2 \rightarrow H_2O$

#### Equation Structure

- **reactants**
  - o **elements/molecules on the left side of the equation**
  - o **elements/molecules that react chemically**
- **product**
  - o **elements/molecules on the right side of the equation**
  - o **elements/molecules that are produced by a chemical reaction**

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