

Temperature

°C, K

measure of average KE
of motion of particles

Heat

kJ, kcal (Cal) [1 kcal = 4.184 kJ]

measure of total energy transferred from
an object of high E → low E.

Note: A change in T is accompanied by a transfer of heat energy.

Specific heat (c or c_p):

(specific for each substance)

amount of energy required to raise the temperature of 1 g of that substance by 1°C.

change in heat = specific heat x mass x change in T

$$q = c_p \times m \times \Delta T$$

Units: heat absorbed or released, in J;

ΔT in °C or K; m in g; c in J/g°C

Molar heat capacity = heat capacity in J/mol°C instead of J/g°C

Don't forget: the size of 1°C = 1 K, so K can be used in this calculation in place of °C.

Examples

- How much heat energy is released to your body when a cup of hot tea containing 200. g of water is cooled from 65.0°C to body temperature, 37.0°C? (c_p of H₂O = 4.18 J/g°C)
- A 4.50-g nugget of pure gold absorbed 276 J of heat. What was the final temperature of the gold if the initial temperature was 25.0°C? (c_p of gold = 0.129 J/g°C)
- A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance?

Chapter 6 Thermochemistry Test

Hussin A.Rothana



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hundreds of practice questions and thorough answer explanations and proven test taking strategies VETs, Complete Preparation for the Veterinary Entrance Tests, 2000 *Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed* Robert K. Wismer, 1982 *Learning Thermochemistry* Karen Chang, 2009 *General Chemistry* James P. Birk, Donald M. Baird, 1994-03 Chemistry George M. Bodner, Harry L. Pardue, 1989-01-23 A modern experimental approach to first year chemistry This unique introductory account employs experimental observations to construct the principles of general chemistry An early introduction to observable descriptive chemistry lays the basis for the well developed exposition that follows **Kaplan SAT II: Chemistry 2004-2005** Kaplan, 2004-03-02 At head of title Kaplan **General Chemistry with Qualitative Analysis** Ralph H. Petrucci, Robert K. Wismer, 1983

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