

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



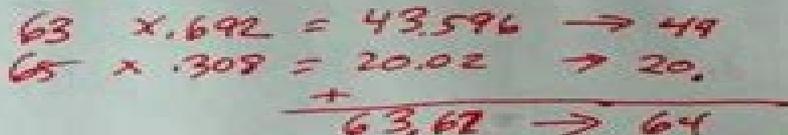
Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Jianjun Gao



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Table of Contents Chemistry 1 Honors Final Exam Study Guide

1. Understanding the eBook Chemistry 1 Honors Final Exam Study Guide
 - The Rise of Digital Reading Chemistry 1 Honors Final Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 1 Honors Final Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 1 Honors Final Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 1 Honors Final Exam Study Guide
 - Personalized Recommendations
 - Chemistry 1 Honors Final Exam Study Guide User Reviews and Ratings

- Chemistry 1 Honors Final Exam Study Guide and Bestseller Lists
- 5. Accessing Chemistry 1 Honors Final Exam Study Guide Free and Paid eBooks
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 - Chemistry 1 Honors Final Exam Study Guide Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Chemistry 1 Honors Final Exam Study Guide Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Chemistry 1 Honors Final Exam Study Guide
 - Highlighting and Note-Taking Chemistry 1 Honors Final Exam Study Guide
 - Interactive Elements Chemistry 1 Honors Final Exam Study Guide
- 8. Staying Engaged with Chemistry 1 Honors Final Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 1 Honors Final Exam Study Guide
- 9. Balancing eBooks and Physical Books Chemistry 1 Honors Final Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 1 Honors Final Exam Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 1 Honors Final Exam Study Guide
 - Setting Reading Goals Chemistry 1 Honors Final Exam Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 1 Honors Final Exam Study Guide
 - Fact-Checking eBook Content of Chemistry 1 Honors Final Exam Study Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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