

Chemistry - Chapter 7 Study Guide

Section 7.1:

- *Valence electrons* – electrons in the outer energy level
 - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
 - Look at the group number
 - Group 1A = 1 valence electron
 - Group 7A = 7 valence electrons
 - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
 - Written using the following steps:
 1. Write the element symbol – representing the nucleus and inner electrons
 2. Put one dot for each valence electrons,
 3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
 - Lose electrons to attain a noble gas configuration
 - Written like Na^{1+}
 - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
 - Gain electrons to attain a noble gas configuration
 - Written like P^{3-}
 - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
 - Also called salts
 - Made from a cation and an anion
 - Which means they're made of a metal and a non-metal
- ***if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
 - Written like Ca_3P_2

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J Dewey



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1. Understanding the eBook Chemistry Chapter 7 Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 7 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 7 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 7 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chapter 7 Study Guide Answers User Reviews and Ratings
 - Chemistry Chapter 7 Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Chapter 7 Study Guide Answers Free and Paid eBooks
 - Chemistry Chapter 7 Study Guide Answers Public Domain eBooks
 - Chemistry Chapter 7 Study Guide Answers eBook Subscription Services
 - Chemistry Chapter 7 Study Guide Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 7 Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 7 Study Guide Answers Compatibility with Devices
 - Chemistry Chapter 7 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 7 Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chapter 7 Study Guide Answers
 - Interactive Elements Chemistry Chapter 7 Study Guide Answers

8. Staying Engaged with Chemistry Chapter 7 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 7 Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 7 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 7 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 7 Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 7 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 7 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 7 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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