



## 6: Equilibrium

Chemistry Worksheet

Standard: 11th Science

Subject: Chemistry

Student Name: \_\_\_\_\_

Q1. Which of the following is not a buffer solution?

1.  $\text{CH}_3\text{COOH} + \text{CH}_3\text{COONa}$
2.  $\text{H}_2\text{BO}_3 + \text{Na}_2\text{BO}_3$
3.  $\text{HClO}_4 + \text{NaClO}_4$
4.  $\text{NH}_4\text{OH} + (\text{NH}_4)_2\text{SO}_4$

Q2. 0.10M  $\text{CH}_3\text{COOH}$  is 1.34% ionised, calculate its  $K_a$ :

1.  $1.8 \times 10^{-1}$
2.  $1.8 \times 10^{-4}$
3.  $5 \times 10^{-4}$
4.  $4 \times 10^{-5}$

Q3. Fear or excitement generally causes one to breathe rapidly and it results in the decrease of concentration of  $\text{CO}_2$  in blood.

In what way, it will change the pH of blood?

1. pH will increase
2. pH will decrease
3. No change
4. pH will be 7

Q4. Pure ammonia is placed in a vessel at a temperature where its dissociation constant ( $\alpha$ ) is appreciable. At equilibrium :

1.  $K_p$  does not change significantly with pressure.
2.  $\alpha$  does not change with pressure.
3. The concentration of  $\text{NH}_3$  does not change with pressure.
4. The concentration of hydrogen is less than that of nitrogen.

Q5. Which of the following salts will give highest pH in water?

1. KCl
2. NaCl
3.  $\text{Na}_2\text{CO}_3$
4.  $\text{CuSO}_4$

Q6. "An acid is a substance that is capable of donating a hydrogen ion  $\text{H}^+$  and bases are substances capable of accepting a



# Chemistry Chemical Equilibrium Test Answers

**Claire Aldridge, Karl Lee, Kaplan, Inc**



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