

QUESTION 1

Volume of pipette used 25.00 cm^3
 (V₀)

Titration	Rough	1	2	3
Final burette reading (cm ³)	20.80	41.40	20.40	40.90
Initial burette reading (cm ³)	0.00	20.80	0.00	20.40
Volume of A used	20.80	20.60	20.40	20.50

② ^{average} Volume of A used = $\frac{20.60 + 20.40 + 20.50}{3}$
 $= 20.50 \text{ cm}^3$

③(i) Concentration of A in mol/dm³ = $\frac{\text{conc. in g/dm}^3}{\text{molar mass}}$
 $= \frac{4.4 \text{ g/dm}^3}{2[13] + 32 + 4[16]}$
 $= \frac{4.4}{98} = 0.0449 \text{ mol/dm}^3$

③(ii) Concentration of B in mol/dm³, C_B

Using $\frac{C_A V_A}{C_B V_B} = \frac{n_A}{n_B}$

$$\frac{0.0449 \times 20.50}{C_B \times 25} = \frac{1}{1}$$

$$C_B = \frac{0.0449 \times 20.50}{25} = 0.03287 \text{ mol/dm}^3$$

Chemistry Practical Question And Answer

Camilla Rothe



Chemistry Practical Question And Answer:

Whispering the Strategies of Language: An Psychological Quest through **Chemistry Practical Question And Answer**

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