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fewer molecules  $\text{N}_2$  g  $3\text{H}_2$  g  $2\text{NH}_3$  g for every two molecules of ammonia made four molecules of reactant are used up this equilibrium shifts to the right with an increase in pressure

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web answer in a chemical reaction chemical equilibrium is the state in which the forward reaction rate and the reverse reaction rate are equal the result of this equilibrium is that the concentrations of the reactants and the products do not change

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web if the rate of the forward reaction decreases increases stays the same more products are formed in a chemical reaction at equilibrium the rate of the forward reaction is equal to the rate of the reverse reaction if the rate of the forward reaction increases more products are formed

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web in order to physically measure the rate of consumption of the bisulfite ion we will look for a color change in the reaction vessel the solution will change from clear to dark blue if you have the following equilibrium reaction  $\text{CO}_2$  g  $2\text{H}_2$  g  $\text{CH}_3\text{OH}$  g

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since at least the 17th century with notable practitioners in previous centuries see the article history

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