

Quality Control Interview Questions & Answers

Redox titrations: Where the titrant is oxidizing agents or reducing agents.

Q47. What is the Ultraviolet (UV) and visible spectroscopy range?

The range of UV Spectroscopy is 200-400 nm, and visible spectroscopy ranges from 400- 800 nm.

Q48. What is the use of UV Spectroscopy?

Spectroscopy can be used to detect functional groups, and impurities, and perform qualitative and quantitative analyses.

Q49. What is meant by the solution?

A solution is a mixture of liquids, gases, and solids, the solution consists of many different types of solutes like salts, oxygen, and organic molecules.

Q50. Describe the saturated and unsaturated solutions

A saturated solution is defined as a solution in which a solvent is not capable of dissolving any more solute at a given temperature.

At a given temperature, an unsaturated solution is one in which the solvent is Capable of dissolving any extra solute.

Q51. What is the difference between qualitative and quantitative analysis?

Qualitative analysis involves the identification of the compound or chemical based on their chemical (absorption, emission) or physical properties, e.g., melting point and boiling point.

Quantitative analysis: this involves estimation or determination of the concentration or amount of the chemical compounds or components.

Q52. Explain the principle of ultraviolet spectroscopy?

Ultraviolet spectroscopy uses light in the UV part of the electromagnetic spectrum. UV absorption spectra form when the outer electrons of a molecule or an atom absorb energy and move from a lower to a higher energy level. The wavelength absorbance of each molecule is unique.

Q53. Define Molarity?

A number of moles of solute per litre solution. Molarity is denoted with a capital "M".

Q54. Define molality?

The number of moles of solute per kilogram solvent. It is denoted with a small "m".

Q55. Define normality?

The number of moles equivalent per litre solution.

Q56. What are buffer solutions?

A buffer solution is an aqueous solution consisting of a mixture of a weak acid and its conjugate base or vice versa.

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