



Fig 5.3. Energy Levels

- The first energy level (labelled 1) can hold up to only two electrons.
- The second energy level (labelled 2) can hold a maximum of eight electrons. This energy level is filled after the first energy level and before the third level.
- The third energy level (labelled 3) can hold a maximum of 18 electrons, however when 8 electrons are in the third level there is a degree of stability; and other electrons are added to the fourth energy level.
- Electrons must first fill the first energy level before they start occupying the second energy level.

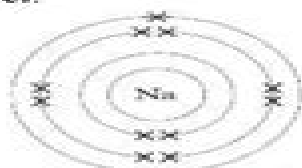


Fig 5.4 Shell model for sodium (Na)

Number of protons = 11, number of electrons = 11

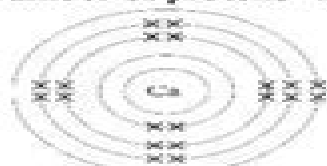


Fig 5.5 Shell model for calcium (Ca)

Number of protons = 20, number of electrons = 20

*Atomic number, mass numbers, sub-atomic particles and electronic configuration of the first 20 elements in the Periodic Table*

# Olevel Intergrated Science Notes

**Ying-Ying Zheng**



## **Olevel Intergrated Science Notes:**

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