



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - CO_2
 - CaCl_2
 - C_2H_6
 - H_2SO_4

- Classify each of the following as either an ionic or a molecular compound.
 - CaCl_2
 - C_2H_6
 - H_2SO_4
 - HCl
 - H_2O
 - KNO_3
 - NaCl
- Write the formula for each of the following molecular compounds.
 - Carbon dioxide
 - Hydrogen peroxide
 - Carbon monoxide
 - Diphosphorus pentoxide
 - Dinitrogen pentoxide
 - Diphosphorus trioxide
 - Dinitrogen trioxide
- Write the name for each of the following molecular compounds.
 - CO_2
 - H_2O_2
 - CO
 - P_2O_5
 - N_2O_5
 - P_2O_3
 - N_2O_3

SECTION 5.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride, HF .
- Show covalent bonding in methane, CH_4 .
- Show covalent bonding in carbon dioxide, CO_2 .



- Show covalent bonding in acetylene, C_2H_2 .
- Draw Lewis structures for each of the following molecular compounds.
 - CO_2
 - H_2O_2
 - CO
 - P_2O_5
 - N_2O_5
 - P_2O_3
 - N_2O_3



SECTION 5.3 BONDING IN THE LITHIUM

- Write the Lewis structure for each of the following molecular compounds.
 - CO_2
 - H_2O_2
 - CO
 - P_2O_5
 - N_2O_5
 - P_2O_3
 - N_2O_3
- Write the Lewis structure for each of the following molecular compounds.
 - CO_2
 - H_2O_2
 - CO
 - P_2O_5
 - N_2O_5
 - P_2O_3
 - N_2O_3

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J Elliott



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