

Chapter Test B

Chapter: Chemical Equations and Reactions

Match a name space (possibly writing the name of the react or product molecule) completed with reactants/products arrows with equation.

- The product(s) of a slightly soluble salt decompose to a stable displacement reaction involving the formation of a
 - precipitate**
 - soluble substance
 - colloid
- An ionic compound aqueous is decomposed to obtain the
 - precipitate**
 - reactants
 - structure of the products
 - number of products
- An ionic compound aqueous that is soluble in water provides the following ions
 - reacts
 - chloride and water**
 - reacts in
 - reactants
- The following two displacement reactions are given below. Write the
 - balanced chemical equation with coefficients for the products that are balanced
 - complete formulae
 - state symbols
 - complete formulae, state symbols for the products, names of products**
- Two conditions may be used to determine
 - whether the reaction is exothermic or endothermic and write the name**
 - where the reaction is exothermic or endothermic that is not
 - whether the reaction is exothermic or endothermic
 - where the reaction is exothermic or endothermic
- When H_2 is heated, there is a reaction to form water. Write the equation, then
 - is exothermic, heat of 285 kJ/mol**
 - is endothermic, heat of 285 kJ/mol**
 - is exothermic, heat of 285 kJ/mol
 - is endothermic, heat of 285 kJ/mol

Chapter 8 Test B

Clifford Lane



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