

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

The periodic table is color-coded by groups and subshells. The groups are numbered 1 through 18. The subshells are labeled as follows:

- s-subshell:** Groups 1 and 2 (pink).
- p-subshell:** Groups 13 through 18 (green).
- d-subshell:** Groups 3 through 10 (blue).
- f-subshell:** Groups 7 through 14 (orange).

A diagram for Hydrogen (H) shows its atomic structure with a nucleus and a single electron shell.



ATOMIC STRUCTURE

Atomic Number, Z , is of protons in the nucleus. For neutral atom, $Z = +$ of electrons. For ions of charge, $+Q$, $Z = +Q$ (cation) or $Z = -Q$ (anion). Atomic mass, A , is the sum of protons and neutrons in the nucleus. Isotopes atoms with the same Z , different A . The atomic mass has a number of protons and neutrons. A number of isotopes is practical with use the atomic mass. Number of Atomic Weight is practical average of atomic masses for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Quantum Numbers:

- Principal Quantum Number (n):** $n = 1, 2, 3, \dots$
- Angular Momentum Quantum Number (l):** $l = 0, 1, 2, \dots, n-1$
- Magnetic Quantum Number (m_l):** $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$
- Spin Quantum Number (m_s):** $m_s = +\frac{1}{2}, -\frac{1}{2}$

Orbitals:

- s-orbital:** $l = 0$, 1 type
- p-orbital:** $l = 1$, 3 types
- d-orbital:** $l = 2$, 5 types
- f-orbital:** $l = 3$, 7 types

NUCLEAR CHEMISTRY

Nuclear reactions alter the nucleus. Particles emitted or absorbed. Alpha particle, α , is ${}^4_2\text{He}$. Beta particle, β , is ${}^0_{-1}\text{e}$. Gamma ray, γ , is high energy electromagnetic radiation. Positron, β^+ , is ${}^0_{+1}\text{e}$. Neutron, n , is ${}^1_0\text{n}$. Proton, p , is ${}^1_1\text{p}$. Electron, e , is ${}^0_{-1}\text{e}$.

Types of processes:

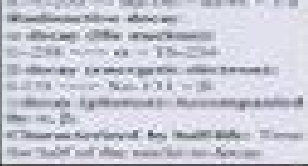
- Radioactive decay:** ${}^A_Z\text{X} \rightarrow {}^A_Z\text{Y} + \text{particle}$
- Transmutation:** ${}^A_Z\text{X} + \text{particle} \rightarrow {}^A_Z\text{Y} + \text{particle}$
- Fission:** ${}^A_Z\text{X} + \text{particle} \rightarrow \text{fission products} + \text{energy}$
- Fusion:** ${}^A_Z\text{X} + {}^A_Z\text{Y} \rightarrow {}^A_Z\text{Z} + \text{energy}$

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe all many-electron atoms and ions. A many electron atom has more electrons than single orbitals and orbital properties. The main point is that orbitals have different energies. The filling of orbitals is guided by the Aufbau Principle.

n	l	m_l	orbitals
1	0	0	1 orbital s-orbital
2	0	0	1 orbital s-orbital
2	1	0, ±1	3 orbitals p-orbitals
3	0	0	1 orbital s-orbital
3	1	0, ±1	3 orbitals p-orbitals
3	2	0, ±1, ±2	5 orbitals d-orbitals
4	0	0	1 orbital s-orbital
4	1	0, ±1	3 orbitals p-orbitals
4	2	0, ±1, ±2	5 orbitals d-orbitals
4	3	0, ±1, ±2, ±3	7 orbitals f-orbitals

Atomic configuration: Shows orbital occupancy. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. An orbital may hold up to two electrons, one with spin up, $+\frac{1}{2}$, and one with spin down, $-\frac{1}{2}$. Hund's Rule: Electrons fill orbitals of equal energy in a way that maximizes the number of unpaired electrons. Aufbau Principle: Orbitals are filled in order of increasing energy. Shielding: Inner electrons shield outer electrons from the nucleus. Penetration: s-orbitals penetrate inner shells more than other orbitals. Exchange energy: Exchange of electrons between orbitals of a chemical bond is a stabilizing factor.



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