

Ideal gas law - problems and solutions

1. Ideal gases in a closed container initially have volume V and temperature T . The final temperature is $5/4T$ and the final pressure is $2P$. What is the final volume of gases.

Known:

Initial volume (V_1) = V

Initial temperature (T_1) = T

Final temperature (T_2) = $5/4 T$

Initial pressure (P_1) = P

Final pressure (P_2) = $2P$

Wanted: Final volume (V_2)

Solution:

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{P V}{T} = \frac{(2P) V_2}{5/4 T}$$

$$\frac{V}{1} = \frac{2 V_2}{5/4}$$

$$5/4 V = 2 V_2$$

$$V_2 = \frac{5V}{4} \cdot \frac{2}{1}$$

$$V_2 = \frac{5V}{4} \times \frac{1}{2}$$

$$V_2 = \frac{5V}{8}$$

2. Determine the volume of 2.00 moles of gases (ideal gas) at STP. STP = Standard Temperature and Pressure.

Known:

Moles of gas (n) = 2 moles

Standard temperature (T) = $0^\circ\text{C} = 0 + 273 = 273$ Kelvin

Standard pressure (P) = 1 atm = 1.013×10^5 Pa

Universal gas constant (R) = 8.315 Joule/mole.Kelvin

Wanted: Volume of gases (V)

Solution:

Read more:

<https://physics.gurumuda.net/ideal-gas-law-problems-and-solutions.htm>

Gas Law Problems With Solutions

L Cohen

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Table of Contents Gas Law Problems With Solutions

1. Understanding the eBook Gas Law Problems With Solutions
 - The Rise of Digital Reading Gas Law Problems With Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Problems With Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Law Problems With Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Problems With Solutions
 - Personalized Recommendations
 - Gas Law Problems With Solutions User Reviews and Ratings

- Gas Law Problems With Solutions and Bestseller Lists
- 5. Accessing Gas Law Problems With Solutions Free and Paid eBooks
 - Gas Law Problems With Solutions Public Domain eBooks
 - Gas Law Problems With Solutions eBook Subscription Services
 - Gas Law Problems With Solutions Budget-Friendly Options
- 6. Navigating Gas Law Problems With Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Law Problems With Solutions Compatibility with Devices
 - Gas Law Problems With Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Law Problems With Solutions
 - Highlighting and Note-Taking Gas Law Problems With Solutions
 - Interactive Elements Gas Law Problems With Solutions
- 8. Staying Engaged with Gas Law Problems With Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Law Problems With Solutions
- 9. Balancing eBooks and Physical Books Gas Law Problems With Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Law Problems With Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Law Problems With Solutions
 - Setting Reading Goals Gas Law Problems With Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Law Problems With Solutions
 - Fact-Checking eBook Content of Gas Law Problems With Solutions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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