

Chapter 1
Properties of Matter

Properties of Matter

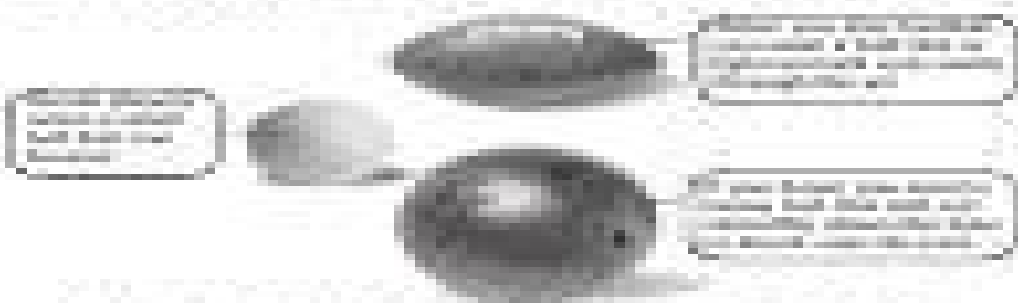
QUESTIONS

Are you able distinguish, how these questions are asked?

- Which one(s) are asked and how they phrased?
- Which are the important and how they phrased?

What are Physical Properties?

Start with the title in the chapter before you start to read the chapter. Focus on how various properties



of different substances are demonstrated. You can see observed without changing the identity of the substance. However the properties of an object change as it reacts. For example, you can identify an object when it changes color, when it melts, evaporates. The properties of an object are determined by the properties of the particles the object consists of.

Many of the physical properties of materials that make up objects can be measured. These properties include density, hardness, magnetism, and the ability to conduct heat and electricity.

Some physical properties depend on how much of a material you have. In other words, a large amount of a material may have different properties than a smaller amount of the material. Other physical properties do not depend on how much of the material you have. They show that no matter how large or small something is, it has the same physical property or always the same

LEARNING OBJECTIVE

Identify the physical and chemical properties of substances. Distinguish between physical and chemical properties. Measure mass and volume.

LEARNING OBJECTIVE

Identify the physical and chemical properties of substances.

LEARNING OBJECTIVE

Identify the physical and chemical properties of substances.

Chapter 2 Properties Of Matter Wordwise Answer Key

George Graybill



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web answer the following questions based on the solubility curve below which salt is least soluble in water at 20 c 2 how many grams of potassium chloride can be dissolved in 200 g of water at 80 c io 3 at 40 c how much potassium nitrate coin be dissoiu tl n 30d g of water w 4 which salt shows the least change

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web chemistry solubility curves virtual lab curves to clearly show the solubility of the four ionic compounds in the data table below the following tips will help you make the best possible graph make the graph large it should fill three fourths of the graph paper

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web solvents the larger volume of the mixture and solutes the smaller volume of the mixture for example a hot chocolate is a solution in which the solute the chocolate powder is dissolved in the solvent the milk or water the solute and solvent can be

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web they should be able to come up with the following answers stir the water increase the temperature and increase the amount of water alternately you may challenge the students to a contest

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web use a solubility curve to determine whether a solution is saturated or unsaturated and to estimate the amount of excess solute that remains undissolved in the solution

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web give results of all calculations in blue cells you must put your answers in the indicated cells as the grade book pulls your data from those cells do not put any units in the number cells use the unit boxes given graph tab transfer from data page the temp and concentration of salt in units of g salt 100 g water

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web 30 09 99 solubility curves class data write down the data answer questions on bottom of the sheet results of lab solubility of KClO_3 unsaturated saturation and solubility curves in this experiment we expressed concentration in units of g solute g H_2O putting concentrations in g solute g solution would be tedious and impractical see fig 2

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web use the solubility curve to answer these questions determine how much potassium nitrate will dissolve in 20 g of water at 40 c 200 cm³ of saturated lead ii nitrate solution was prepared at a temperature of 90 c

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web sample answer sugar has a much greater solubility in hot water than cold water so its solubility curve probably resembles the sloping solubility curve of potassium nitrate temperature solubility g 100 ml 10 c 36 g 100 ml 20 c 36 g 100 ml 30 c 37 g 100 ml 40 c 37 g 100 ml 50 c 38 g 100 ml 60 c 38 g 100 ml 70 c 39 g 100 ml 80

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