

Mathematically

$$\frac{P^0 - P^s}{P^0} = x_2$$

Here P^0 = Pressure of solvent in pure stage.

P^s = Pressure of solution

x_2 = Mole fraction of solute.

DERIVATION OF THE LAW

Pressure of solvent in solution

= Mole fraction of solvent \times pressure of solvent in pure stage

If solute is non-volatile

Total pressure of solution = Pressure due to solvent

= Mole fraction of solvent \times pressure of solvent in pure stage

So $P_{\text{solvent}} = x_1 \cdot P_{\text{solution}}$ Here x_1 is mole fraction of solvent.

$$\frac{P_{\text{solvent}}}{P_{\text{solution}}} = x_1$$

Subtract each term from 1 on both side.

$$1 - \frac{P_{\text{solvent}}}{P_{\text{solution}}} = 1 - x_1$$

$$\frac{P_{\text{solution}} - P_{\text{solvent}}}{P_{\text{solution}}} = x_2$$

2.11 IDEAL & NON-IDEAL SOLUTION

Sr.	Ideal solution	Non-Ideal solution
1.	A solution which obey Roult's law is called ideal solution.	A solution which do not obey Roult's law is called non-ideal solution.
2.	Interaction between components of mixture i.e.-between A&B is some as found in A-A & B-B molecule.	Interaction between components of mixture i.e.-between A&B is not some as found in A-A & B-B molecule.

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L Darling-Hammond



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Table of Contents Class 12th Notes Chemistry Chapter Wise Notes

1. Understanding the eBook Class 12th Notes Chemistry Chapter Wise Notes
 - The Rise of Digital Reading Class 12th Notes Chemistry Chapter Wise Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Class 12th Notes Chemistry Chapter Wise Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Class 12th Notes Chemistry Chapter Wise Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Class 12th Notes Chemistry Chapter Wise Notes
 - Personalized Recommendations

- Class 12th Notes Chemistry Chapter Wise Notes User Reviews and Ratings
- Class 12th Notes Chemistry Chapter Wise Notes and Bestseller Lists
- 5. Accessing Class 12th Notes Chemistry Chapter Wise Notes Free and Paid eBooks
 - Class 12th Notes Chemistry Chapter Wise Notes Public Domain eBooks
 - Class 12th Notes Chemistry Chapter Wise Notes eBook Subscription Services
 - Class 12th Notes Chemistry Chapter Wise Notes Budget-Friendly Options
- 6. Navigating Class 12th Notes Chemistry Chapter Wise Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Class 12th Notes Chemistry Chapter Wise Notes Compatibility with Devices
 - Class 12th Notes Chemistry Chapter Wise Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Class 12th Notes Chemistry Chapter Wise Notes
 - Highlighting and Note-Taking Class 12th Notes Chemistry Chapter Wise Notes
 - Interactive Elements Class 12th Notes Chemistry Chapter Wise Notes
- 8. Staying Engaged with Class 12th Notes Chemistry Chapter Wise Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Class 12th Notes Chemistry Chapter Wise Notes
- 9. Balancing eBooks and Physical Books Class 12th Notes Chemistry Chapter Wise Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Class 12th Notes Chemistry Chapter Wise Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Class 12th Notes Chemistry Chapter Wise Notes
 - Setting Reading Goals Class 12th Notes Chemistry Chapter Wise Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Class 12th Notes Chemistry Chapter Wise Notes
 - Fact-Checking eBook Content of Class 12th Notes Chemistry Chapter Wise Notes

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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