

# PHYSICAL PHARMACY

**PHYSICAL CHEMISTRY** – branch of chemistry concerned with physical properties and structure of matter and with laws and theories of physical and chemical changes

**PHYSICAL PHARMACY** – application of physical chemistry in pharmacy; study of physicochemical properties of substances used in drug formulation

## I. FORCES OF ATTRACTION

### A. INTRAMOLECULAR – intra (within) the molecule

- **Ionic Bond** – transfer of electrons; metal (electron donor) + non-metal (electron acceptor); (ex. Na:  $p^+ = 11$ ;  $e^- = 11 \rightarrow p^+ = 11$ ;  $e^- = 10 \rightarrow$  charge = +1)

#### ○ PERIODIC TRENDS REPRESENTING BONDS

- **Ionization Energy** – energy required remove electron (low M; high NM)
- **Electronegativity** – ability to attract electrons (low M; high NM)

- **Covalent Bonds** – sharing of electrons; non-metal + non-metal; polarity – difference in electronegativity (symmetrical = non-polar)

- **Polar** – unequal sharing of electrons; dipole moment – (ex.  $H^{\delta+}-Br^{\delta-}$ )
- **Non Polar** – equal sharing of electrons (ex.  $Br-Br$ )

### B. INTERMOLECULAR – inter (between) the molecules

- **Van der Waals**

- **Dipole-Dipole**  $\rightarrow P + P$ ; Keesom Forces; orientation effect
- **Dipole-Induced Dipole**  $\rightarrow P + NP$ ; Debye Forces; inductive effect
- **Induced Dipole-Induced Dipole**  $\rightarrow NP + NP$ ; London Forces; dispersive effect

- **Ion-Induced Dipole**  $\rightarrow ion + NP$

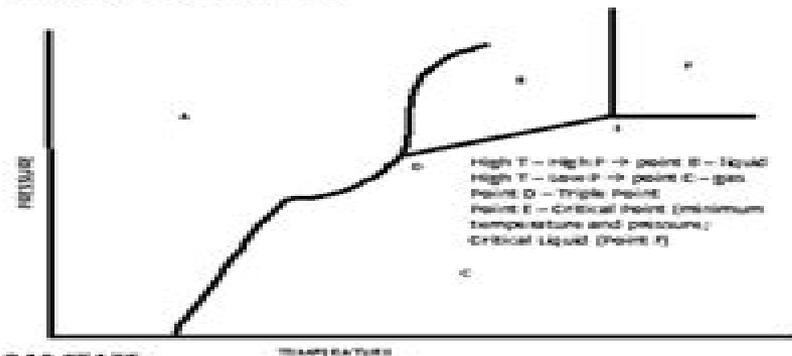
$I_2 \rightarrow NP$  (solubility = 1:2960)

$I_2 + NaI/KI$  (solubilizing agent)  $\rightarrow I_2 + I_3^-$  = increase solubility

- **Hydrogen Bond** – hydrogen atom + electronegative atom (N, O, F)



## PHASE DIAGRAM



### GAS STATE:

- **Kinetic Molecular Theory**

- Gases are composed of small particles called atoms; total volume is negligible in relation to space where they are confined
- Particles do not attract one another but move with complete independence
- Particles exhibit continuous random motion owing to their kinetic energy
- Gases exhibit perfect elasticity (no energy exchange)

## II. GAS LAWS

- at standard temperature (Kelvin =  $^{\circ}C + 273$ ) and pressure (1 atm = 760 mmHg or torr; 760 mmHg;  $1.01325 \times 10^5$  dynes/cm<sup>2</sup>;  $1.01325 \times 10^5$  Pa or N/m<sup>2</sup>; 0.1325 bar)

- **Boyle's Law** – constant temperature; increase P = decrease V; inversely related

$$P \propto \frac{1}{V}$$

- \* density = increase altitude; decrease volume;  $P_1V_1 = P_2V_2$

- **Charles' Law** – constant pressure; increase T = increase V

$$T \propto V; \frac{T_1}{T_2} = \frac{V_1}{V_2} \text{ or } \frac{T_1}{V_1} = \frac{T_2}{V_2}$$

- \*  $H_2O$  – 1.0g/ml  $\rightarrow 4^{\circ}C$

- **Gay Lussac's Law**  $\rightarrow$  constant volume

$$P \propto T; \frac{P_1}{P_2} = \frac{T_1}{T_2} \text{ or } \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- \*  $121^{\circ}C$  – 15-20psi

# Physical Pharmacy Lecture Notes

**Michael Brown**



## **Physical Pharmacy Lecture Notes:**

## Embracing the Melody of Term: An Psychological Symphony within **Physical Pharmacy Lecture Notes**

In a world eaten by screens and the ceaseless chatter of immediate transmission, the melodic splendor and psychological symphony created by the written term frequently diminish into the background, eclipsed by the persistent noise and disturbances that permeate our lives. However, nestled within the pages of **Physical Pharmacy Lecture Notes** a wonderful fictional treasure full of fresh thoughts, lies an immersive symphony waiting to be embraced. Constructed by an elegant composer of language, that fascinating masterpiece conducts visitors on a psychological journey, well unraveling the concealed songs and profound impact resonating within each cautiously constructed phrase. Within the depths of the moving examination, we will discover the book is key harmonies, analyze their enthralling publishing model, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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