

## Determination of $K_a$ of a Weak Acid

(Show all data and calculations in your lab notebook)

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Concentration of your standardized base       $0.090 \pm 0.0021$

**Table 1: Experimental Data**

	Trial 1	Trial 2	Trial 3
Volume of unknown weak acid used, mL	34 mL	32 mL	33 mL
Volume of base required to reach equivalence point, mL	32.3 mL	30.2 mL	31.4 mL

**Table 2: Experimental Results**

	Trial 1	Trial 2	Trial 3	Mean $\pm$ SD
Concentration of unknown weak acid, M	0.0846	0.0840	0.0847	0.084 $\pm$ 0.00037
Volume of base required to reach half-equivalence point, mL	16.2	15.1	15.7	
pH at the half-equivalence point (from graph)	4.62	4.58	4.53	
$pK_a$ of the unknown weak acid	4.62	4.58	4.53	
$K_a$ of the unknown weak acid	$2.3 \times 10^{-5}$	$2.6 \times 10^{-5}$	$2.9 \times 10^{-5}$	$2.6 \times 10^{-5} \pm 3 \times 10^{-6}$

# Determination Of Ka Lab Report Answers

**Xiang Xie**



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