

QUESTION 1

Volume of pipette used 25.00 cm^3
 (V₀)

| Titration | Rough | 1 | 2 | 3 |
|--|-------|-------|-------|-------|
| Final burette reading (cm ³) | 20.80 | 41.40 | 20.40 | 40.90 |
| Initial burette reading (cm ³) | 0.00 | 20.80 | 0.00 | 20.40 |
| Volume of A used | 20.80 | 20.60 | 20.40 | 20.50 |

② ^{average} Volume of A used = $\frac{20.60 + 20.40 + 20.50}{3}$
 $= 20.50 \text{ cm}^3$

③(i) Concentration of A in mol/dm³ = $\frac{\text{conc. in g/dm}^3}{\text{molar mass}}$
 $= \frac{4.4 \text{ g/dm}^3}{2[13] + 32 + 4[16]}$
 $= \frac{4.4}{98} = 0.0449 \text{ mol/dm}^3$

③(ii) Concentration of B in mol/dm³, C_B

Using $\frac{C_A V_A}{C_B V_B} = \frac{n_A}{n_B}$

$$\frac{0.0449 \times 20.50}{C_B \times 25} = \frac{1}{1}$$

$$C_B = \frac{0.0449 \times 20.50}{25} = 0.03287 \text{ mol/dm}^3$$

Chemistry Practical Question And Answer

Michael Seilmaier



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