

Campbell Biology, 11e (Urry)
Chapter 3 Water and Life

3.1 Multiple-Choice Questions

1) In a single molecule of water, two hydrogen atoms are bonded to a single oxygen atom by _____.

- A) hydrogen bonds
- B) nonpolar covalent bonds
- C) polar covalent bonds
- D) ionic bonds

Answer: C

Bloom's Taxonomy: Knowledge/Comprehension

Section: 3.1

2) The partial negative charge at one end of a water molecule is attracted to the partial positive charge of another water molecule. What is this attraction called?

- A) a covalent bond
- B) a hydrogen bond
- C) an ionic bond
- D) a van der Waals interaction

Answer: B

Bloom's Taxonomy: Knowledge/Comprehension

Section: 3.1

3) The partial negative charge in a molecule of water occurs because _____.

- A) the oxygen atom donates an electron to each of the hydrogen atoms
- B) the electrons shared between the oxygen and hydrogen atoms spend more time around the oxygen atom nucleus than around the hydrogen atom nucleus
- C) the oxygen atom has two pairs of electrons in its valence shell that are not neutralized by hydrogen atoms
- D) one of the hydrogen atoms donates an electron to the oxygen atom

Answer: B

Bloom's Taxonomy: Knowledge/Comprehension

Section: 3.1

4) Sulfur is in the same column of the periodic table as oxygen, but has electronegativity similar to carbon. Compared to water molecules, molecules of H₂S will _____.

- A) have greater cohesion to other molecules of H₂S
- B) have a greater tendency to form hydrogen bonds with each other
- C) have a higher capacity to absorb heat for the same change in temperature
- D) not form hydrogen bonds with each other

Answer: D

Bloom's Taxonomy: Synthesis/Evaluation

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Michal Rosen-Zvi



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1. Understanding the eBook Campbell Biology Chapter 3 Test Bank
 - The Rise of Digital Reading Campbell Biology Chapter 3 Test Bank
 - Advantages of eBooks Over Traditional Books
2. Identifying Campbell Biology Chapter 3 Test Bank
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Campbell Biology Chapter 3 Test Bank
 - User-Friendly Interface
4. Exploring eBook Recommendations from Campbell Biology Chapter 3 Test Bank
 - Personalized Recommendations
 - Campbell Biology Chapter 3 Test Bank User Reviews and Ratings
 - Campbell Biology Chapter 3 Test Bank and Bestseller Lists
5. Accessing Campbell Biology Chapter 3 Test Bank Free and Paid eBooks
 - Campbell Biology Chapter 3 Test Bank Public Domain eBooks
 - Campbell Biology Chapter 3 Test Bank eBook Subscription Services

- Campbell Biology Chapter 3 Test Bank Budget-Friendly Options
- 6. Navigating Campbell Biology Chapter 3 Test Bank eBook Formats
 - ePub, PDF, MOBI, and More
 - Campbell Biology Chapter 3 Test Bank Compatibility with Devices
 - Campbell Biology Chapter 3 Test Bank Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Campbell Biology Chapter 3 Test Bank
 - Highlighting and Note-Taking Campbell Biology Chapter 3 Test Bank
 - Interactive Elements Campbell Biology Chapter 3 Test Bank
- 8. Staying Engaged with Campbell Biology Chapter 3 Test Bank
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Campbell Biology Chapter 3 Test Bank
- 9. Balancing eBooks and Physical Books Campbell Biology Chapter 3 Test Bank
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Campbell Biology Chapter 3 Test Bank
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Campbell Biology Chapter 3 Test Bank
 - Setting Reading Goals Campbell Biology Chapter 3 Test Bank
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Campbell Biology Chapter 3 Test Bank
 - Fact-Checking eBook Content of Campbell Biology Chapter 3 Test Bank
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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