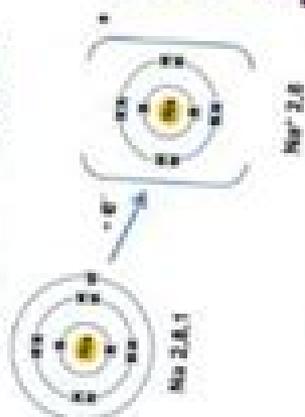


Group 1: The Alkali Metals

All have **one electron** in their **outermost shell**. React by **losing** this **electron** to form a **positive metal ion** with a **1+ charge**.

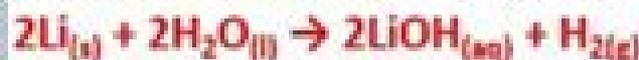


Reactivity increases as you go down the group. This is because:

- **Outermost electron** is in a shell which is **further from the nucleus** so it is **less strongly attracted** and **lost more easily**.
- **Increased shielding** as **more inner shell electrons** so is **less strongly attracted**.

| | | | |
|----|----|--------|--------------------------------------|
| 3 | Li | 6.941 | LEAST REACTIVE ↓ MOST REACTIVE |
| 11 | Na | 22.990 | |
| 19 | K | 39.100 | |
| 37 | Rb | 85.467 | |
| 55 | Cs | 132.9 | |
| 87 | Fr | 223 | |

All react with water to form an **alkaline** solution of **metal hydroxide** and **hydrogen gas**.



You will often be asked what you can **observe**:

- **The metal floats**
- **Fizzing / A gas is produced**
- **The metal 'disappears'**
- **The reactions become more vigorous as you do down the group**

We call them alkali metals because they react with water to form **alkaline solutions** of **hydroxide ions (OH⁻)**. **Universal indicator** will turn **purple**. These solutions are highly **corrosive**



Alkali metals produce characteristic **colours in a flame**. We can use these to identify them:

- **Lithium – crimson**
- **Sodium – yellow**
- **Potassium - lilac**



As you do down group 1:

- **Melting point decreases**
 - They **become softer** / easier to cut.
- Increased distance** and **shielding** makes the **metallic bonding weaker**.



Alkali metals react with oxygen to form **metal oxides**. Metal oxides are **bases** that react with **acids** and **neutralise** them.



Aqa Gcse Chemistry Revision Notes

**Christine Horbury, Andrew
Catterall, Lynn Henfield**



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