

Chemistry - Chapter 7 Study Guide

Section 7.1:

- *Valence electrons* – electrons in the outer energy level
 - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
 - Look at the group number
 - Group 1A = 1 valence electron
 - Group 7A = 7 valence electrons
 - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
 - Written using the following steps:
 1. Write the element symbol – representing the nucleus and inner electrons
 2. Put one dot for each valence electrons,
 3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
 - Lose electrons to attain a noble gas configuration
 - Written like Na^{1+}
 - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
 - Gain electrons to attain a noble gas configuration
 - Written like P^{3-}
 - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
 - Also called salts
 - Made from a cation and an anion
 - Which means they're made of a metal and a non-metal
- ***if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
 - Written like Ca_3P_2

Chemistry Matter And Change Chapter 7 Study Guide Answer Key

K Morrison



Chemistry Matter And Change Chapter 7 Study Guide Answer Key:

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With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them

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