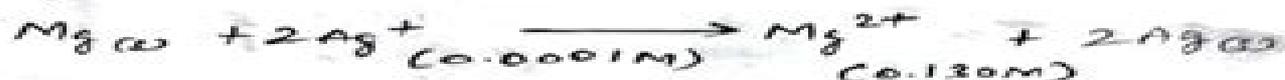


ELECTROCHEMISTRY NUMERICALS:

EXAMPLE PROBLEMS:

Prob: 3.1 Represent the cell in which the following reaction takes place



Calculate its E_{cell} if $E_{\text{cell}}^{\circ} = 3.17\text{V}$

Soln:

cell notation:



$$\begin{aligned} E_{\text{cell}} &= E_{\text{cell}}^{\circ} - \frac{0.0591\text{V}}{2} \log \frac{0.130}{(0.0001)^2} \\ &= 3.17\text{V} - 0.02955 \log \frac{0.13}{(10^{-4})^2} \\ &= 3.17 - 0.02955 \log \frac{0.13}{10^{-8}} \\ &= 3.17 - 0.02955 \log (0.13 \times 10^8) \\ &= 3.17 - 0.02955 \log (13 \times 10^6) \\ &= 3.17 - 0.02955 \log 13 + \log 10^6 \\ &= 3.17 - 0.02955 (1.1139 + 6) \\ &= 3.17 - (0.02955 \times 7.1139) \\ \therefore E_{\text{cell}} &= 3.17 - 0.21 = 2.96\text{V} \checkmark \end{aligned}$$

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**Eugene Franz Roeber, Howard Coon
Parmelee**



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