

**Honors Chemistry**  
**Extra Stoichiometry Problems**

1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate.  
a. Write and balance the chemical equation.



- b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many representative particles (and what type) of silver chloride are produced?

$$\frac{39.02 \text{ g BaCl}_2}{1} \times \frac{1 \text{ mole BaCl}_2}{208.233 \text{ g BaCl}_2} \times \frac{2 \text{ moles AgCl}}{1 \text{ mole BaCl}_2} \times \frac{6.02 \times 10^{23} \text{ f. u. AgCl}}{1 \text{ mole AgCl}} = 2.256 \times 10^{23} \text{ f. u. AgCl}$$

- c. If 410.8 grams of barium nitrate are produced how many grams of silver nitrate were reacted?

$$\frac{410.8 \text{ g Ba}(\text{NO}_3)_2}{1} \times \frac{1 \text{ mole Ba}(\text{NO}_3)_2}{261.335 \text{ g Ba}(\text{NO}_3)_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ mole Ba}(\text{NO}_3)_2} \times \frac{169.872 \text{ g AgNO}_3}{1 \text{ mole AgNO}_3} = 534.1 \text{ g AgNO}_3$$

2. Nitrogen gas is reacted with hydrogen gas to form nitrogen trihydride.

- a. Write and balance the chemical equation.



- b. How many liters of nitrogen trihydride are produced at STP if 80.28 grams of hydrogen gas are reacted in an excess of nitrogen?

$$\frac{80.28 \text{ g H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.016 \text{ g H}_2} \times \frac{2 \text{ moles NH}_3}{3 \text{ moles H}_2} \times \frac{22.4 \text{ L NH}_3}{1 \text{ mole NH}_3} = 594.7 \text{ L NH}_3$$

- c. How many grams of hydrogen are needed to fully react 621.9 grams of nitrogen gas?

$$\frac{621.9 \text{ g N}_2}{1} \times \frac{1 \text{ mole N}_2}{28.014 \text{ g N}_2} \times \frac{3 \text{ moles H}_2}{1 \text{ mole N}_2} \times \frac{2.016 \text{ g H}_2}{1 \text{ mole H}_2} = 134.3 \text{ g H}_2$$

- d. How many liters of nitrogen gas are needed to completely react 90.38 L of hydrogen gas at STP?

$$\frac{90.28 \text{ L H}_2}{1} \times \frac{1 \text{ mole H}_2}{22.4 \text{ L H}_2} \times \frac{1 \text{ mole N}_2}{3 \text{ mole H}_2} \times \frac{22.4 \text{ L N}_2}{1 \text{ mole N}_2} = 30.09 \text{ L N}_2$$

$$\frac{90.28 \text{ L H}_2}{1} \times \frac{\text{OR}}{3 \text{ L H}_2} = 30.09 \text{ L N}_2$$

# Stoichiometry 8b Extra Practice Problems Answers

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Waterman, Prentice-Hall Staff**

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Greenblatt, Sander van Smaalen, 2012-12-06 The field of low dimensional conductors has been very active for more than twenty years It has grown continuously and both the inorganic and organic materials have remarkable properties such as charge and spin density waves and superconductivity The discovery of superconductivity at high temperature in copper based quasi two dimensional conducting oxides nearly ten years ago has further enlarged the field and stimulated new research on inorganic conductors It was obviously impossible to cover such a broad field in a ten day Institute and it seemed pertinent to concentrate on inorganic conductors excluding the high T<sub>c</sub> superconducting oxides In this context it was highly desirable to include both physics and chemistry in the same Institute in order to tighten or in some cases to establish links between physicists and chemists This Advanced Study Institute is the continuation of a series of similar ones which have taken place every few years since 1974 73 participants coming from 13 countries have taken part in this School at the beautiful site of the Centre de Physique des Houches in the Mont Blanc mountain range The scientific programme included

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