

SMALL-SCALE LAB: Making a Solution

LABORATORY RECORD SHEET

Use with SECTION 18.2

SAFETY



Wear safety glasses and follow the standard safety precautions outlined on page 7 of this manual.

PURPOSE

To make a solution and use carefully measured data to calculate the solution's concentration in various units.

MATERIALS

- solid NaCl
- water
- 50-mL plastic volumetric bottle
- balance

PROCEDURE

Weigh a clean, dry, plastic volumetric bottle. Add solid NaCl to approximately the 5-mL mark on the bottle. Weigh the bottle again. Half fill the bottle with water and shake it gently until all the NaCl dissolves. Fill the bottle with water to the 50-mL mark and weigh it again.

ANALYSIS

Using your experimental data, record the answers to the following questions.

1. Percent by mass tells how many grams of solute are present in 100 g of solution.

$$\% \text{ by mass} = \frac{\text{mass of solute}}{\text{mass of solute} + \text{solvent}} \times 100$$

- a. Calculate the mass of the solute (NaCl).
- b. Calculate the mass of the solvent (water).
- c. Calculate the percent by mass of NaCl in the solution.

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**Dana W. Mayo, Ronald M. Pike, David C.
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