

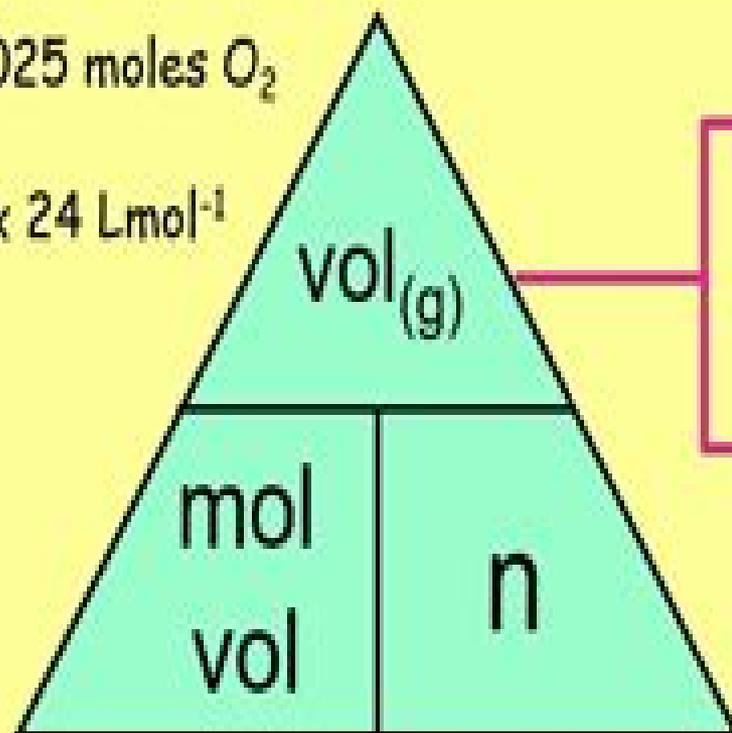
Assume molar volume is 24Lmol^{-1}

Calculate volume of 0.025 moles O_2

$$\text{Vol} = 0.025 \text{ mol} \times 24 \text{ Lmol}^{-1}$$

$$= 0.6 \text{ L}$$

$$= 600 \text{ cm}^3$$



Gas Volume

$$\text{Vol} = n \times \text{mol.vol}$$

Calculations Using Standard Molar Volume

VM Jensen



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Laboratory Experiments for Foundations of Chemistry Ernest R. Toon,1973

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web question 1 what is true of isothermal process a $\Delta t = 0$ b $\Delta u = 0$ c $\Delta q = \Delta w$ d p v constants solution in an isothermal process temperature remains constant $\Delta t = 0$ since internal energy depends on the temperature $\Delta u = 0$ from first law of thermodynamics $\Delta u = \Delta q - \Delta w$ since $\Delta u = 0$ $\Delta q = \Delta w$ also p $v = n$ r T as T is constant p v

