

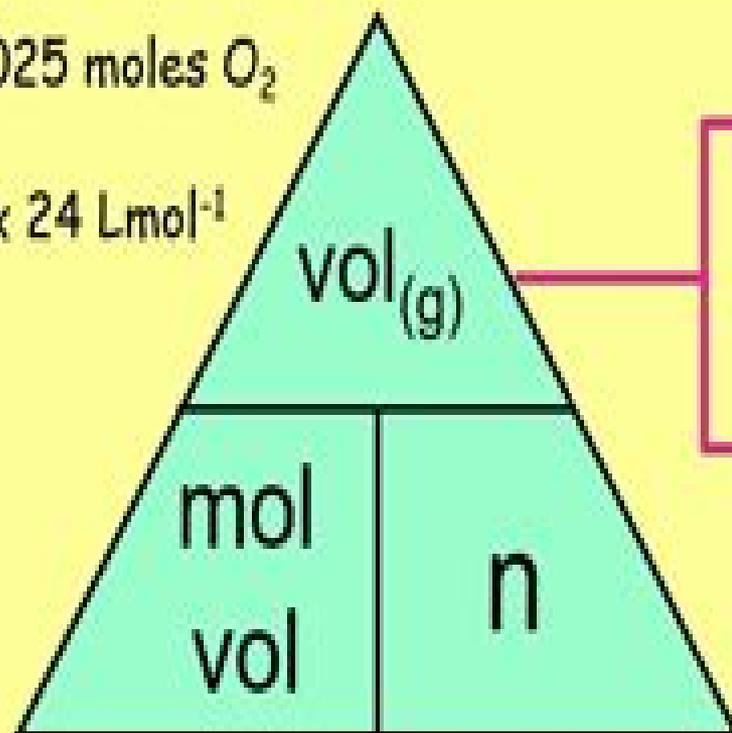
Assume molar volume is 24Lmol^{-1}

Calculate volume of 0.025 moles O_2

$$\text{Vol} = 0.025 \text{ mol} \times 24 \text{ Lmol}^{-1}$$

$$= 0.6 \text{ L}$$

$$= 600 \text{ cm}^3$$



Gas Volume

$$\text{Vol} = n \times \text{mol.vol}$$

Calculations Using Standard Molar Volume

Anthony S. Fauci



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The Calculations of General Chemistry, with Definitions, Explanations, and Problems William Jay Hale, 1920

Fundamentals of Air Sampling Gregory D. Wight, 2018-04-20 There is a growing need for environmental measurement personnel who possess a solid understanding of the techniques of air pollutant sampling This essential book explains the fundamentals of air sampling develops the theory of gas measurement and presents several how to examples of calibration and use of air and gas sampling devices Other topics covered range from the basics of pressure measurement and units conversion to specific discussions regarding the use of a Volatile Organic Sampling Train or a SUMMA polished canister sampling system

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Laboratory Experiments for Foundations of Chemistry Ernest R. Toon,1973

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