11-1 Practice Problems

1. Lead will react with Hydrochloric acid to produce Lead (II) Chloride and Hydrogen gas. How many moles of Hydrochloric acid are needed to completely react with 0.36-mol of lead?

2. How many moles of HNO_3 will be produced when 0.51-mol of N_2O_5 reacts according to the following equation?
$N_2O_5 + H_2O \rightarrow 2HNO_3$

3. Iron will react with Oxygen Gas to produce Fe_2O_3. How many moles of Fe_2O_3 will be produced if 0.18-mol of Fe reacts?

4. How many moles of NaBr will be produced when 0.69-mol of Bromine reacts according to the following equation?
$Br_2 + 2NaI \rightarrow 2NaBr + I_2$

5. Phosphorus will react with Bromine to produce Phosphorus Tribromide. How many moles of Phosphorus Tribromide will be produced if 0.78-mol of Bromine is reacted?

6. Carbon will react with Zinc Oxide to produce Zinc and Carbon Dioxide. How many moles of carbon dioxide will be produced if 0.38-mol of ZnO is completely reacted?

7. How many moles of oxygen will be needed to react with 0.38-mol of C_3H_8 according to the following equation?
$C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

8. Nitrogen gas can react with Hydrogen gas to produce ammonia. How many moles of Nitrogen will be needed to produce 0.48-mol of NH_3?

9. How many moles of hydrogen will be produced if 0.44-mol of CaH_2 reacts according to the following equation?
$CaH_2 + 2H_2O \rightarrow Ca(OH)_2 + 2H_2$

10. How many moles of water will be produced if 2.35-mol of oxygen reacts according to the following equation?
$2C_6H_6 + 15O_2 \rightarrow 12CO_2 + 6H_2O$

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J Spring

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