

**Introduction**

This course is a general chemistry course for students who are interested in pursuing a degree in a health-related field. The course covers the basic principles of chemistry, including atomic structure, chemical bonding, stoichiometry, and thermodynamics. The course is designed to provide students with a solid foundation in chemistry that will be useful in their future careers.

**Prerequisites:** Students must have completed high school chemistry or an equivalent course. Students who do not have the prerequisite course should consult with their advisor for more information.

**Part 1: Atomic Structure**



1. Define the terms atom, element, and compound. Explain the difference between an atom and a molecule. Describe the structure of an atom, including the nucleus and the electron cloud.
2. Explain the periodic table of elements. Describe the trends in atomic size, ionization energy, and electronegativity.
3. Calculate the atomic mass of an element from its isotopic composition.



4. Calculate the mass percent of an element in a compound.
5. Calculate the empirical formula of a compound from its mass percent composition.

| Element | Atomic Mass | Number of Atoms | Mass of Element | Mass of Compound | Mass Percent |
|---------|-------------|-----------------|-----------------|------------------|--------------|
| C       | 12.01       | 1               | 12.01           | 12.01            | 100%         |
| H       | 1.01        | 1               | 1.01            | 13.02            | 7.68%        |
| O       | 16.00       | 1               | 16.00           | 32.01            | 50.31%       |
| N       | 14.01       | 1               | 14.01           | 32.01            | 43.76%       |

6. Calculate the molecular formula of a compound from its empirical formula and molar mass.
7. Calculate the percent yield of a reaction.

**Part 2: Chemical Bonding**



8. Explain the difference between ionic and covalent bonding. Describe the formation of an ionic bond and a covalent bond.
9. Calculate the lattice energy of an ionic compound.
10. Calculate the bond energy of a covalent bond.

$$\Delta H_{\text{rxn}} = \sum \Delta H_{\text{f}}(\text{products}) - \sum \Delta H_{\text{f}}(\text{reactants})$$

11. Calculate the enthalpy change of a reaction.

| Reaction   | Enthalpy Change (kJ/mol) | Equilibrium Constant (K) | Free Energy Change (kJ/mol) | Entropy Change (J/mol·K) |
|--|--------------------------|--------------------------|-----------------------------|--------------------------|
| 1. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$                                    | -92.3                    | > 1                      | -184.6                      | 188.7                    |
| 2. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$                                   | -91.8                    | > 1                      | -163.6                      | 191.6                    |
| 3. $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$   | -1411                    | > 1                      | -282.2                      | 270.2                    |
| 4. $\text{C}_2\text{H}_6 + 7/2\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$ | -1568                    | > 1                      | -313.6                      | 270.2                    |

Work on these problems! Good luck!

# Basic Stoichiometry Phet Lab Homework Exercises Answers

**R Pring**



**Basic Stoichiometry Phet Lab Homework Exercises Answers:**

## Whispering the Secrets of Language: An Psychological Journey through **Basic Stoichiometry Phet Lab Homework Exercises Answers**

In a digitally-driven world wherever screens reign great and immediate conversation drowns out the subtleties of language, the profound techniques and emotional nuances concealed within words frequently move unheard. Yet, nestled within the pages of **Basic Stoichiometry Phet Lab Homework Exercises Answers** a charming fictional value pulsating with natural feelings, lies an exceptional quest waiting to be undertaken. Published by an experienced wordsmith, this charming opus attracts readers on an introspective trip, softly unraveling the veiled truths and profound influence resonating within the very fabric of each word. Within the emotional depths with this moving review, we can embark upon a honest exploration of the book is primary styles, dissect their fascinating publishing model, and succumb to the powerful resonance it evokes strong within the recesses of readers hearts.

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