

REACTIVITY TRENDS IN THE PERIODIC TABLE

Reactivity Trends in the Periodic Table

Reading Passage

Read the passage carefully before answering the questions that follow.

The periodic table organizes elements based on their atomic number and properties, revealing patterns known as periodic trends. One of these trends is reactivity, which describes how likely an element is to participate in chemical reactions. Understanding these trends helps scientists predict how elements will behave in various chemical contexts.

Metals, such as sodium, calcium, and potassium, are highly reactive because they have a single valence electron that they readily lose in chemical reactions. This tendency is due to their large atomic size and weak nuclear energy, making it energetically favorable for them to react with other substances, especially nonmetals. Conversely, noble gases, like helium, neon, and argon, are very inert because they have full valence electron shells. This stable electronic configuration makes them unlikely to form bonds or participate in reactions under normal conditions.

Across periods in the periodic table, reactivity generally decreases as atomic number increases and valence electrons are held more tightly. In addition, atomic size influences reactivity; larger atoms tend to be more reactive if they have fewer valence electrons in their outermost shell. This electronic configuration directly impacts their ability to attract or give up electrons during chemical reactions. The periodic table also shows trends in electronegativity, which is the ability of an atom to attract electrons and thereby reduce its electrical charge. These trends are closely related to reactivity, as elements with high electronegativity (nonmetals) tend to gain electrons, while elements with low electronegativity (metals) tend to lose electrons. Understanding these patterns allows scientists to better predict and control chemical reactions, including developments in material science, industrial chemistry, and environmental systems.

True or False: Reactivity in the Periodic Table

Circle whether each statement is True or False.

- Metals are highly reactive because they have single valence electrons that are difficult to lose.
- Noble gases are very reactive because they have full valence electron shells.
- Reactivity increases across a period, which generally decreases the number of valence electrons.
- Electrons are held more tightly in atoms with smaller atomic numbers than those with more valence electrons.

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DJ Losen



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