

# Class 10th Chemistry Practical Copy

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A solution can be prepared only by combining the solute and solvent. It is a mixture of two substances. The concentration of a solution can be expressed in terms of its mass. It has been expressed with correct number of figures of precision. The following are possible:

1. Mass of solute (g) / Mass of solution (g) × 100  
 2. Mass of solute (g) / Volume of solution (ml) × 100  
 3. Volume of solute (ml) / Volume of solution (ml) × 100

1. Solution having a pH of 7 will be a neutral solution. Solution with pH value above 7 are basic in nature and solution with pH value below 7 are acidic.

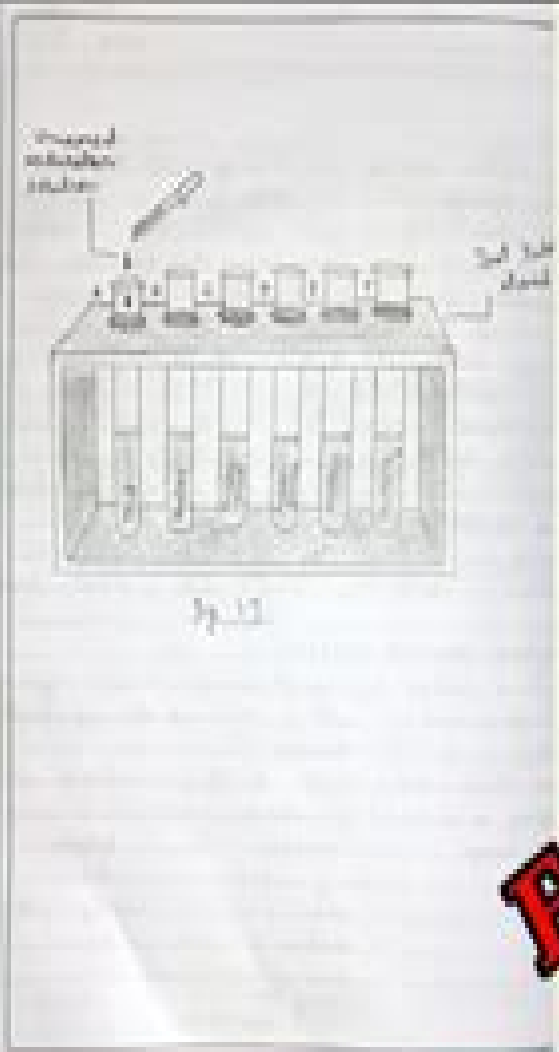
2. pH of a solution can be determined using a pH scale or using an indicator with universal indicator.

3. Universal Indicator Solution

It is a mixture of several indicators having different pH ranges and is used to determine the approximate value of pH of a solution.

4. The colour that indicates the pH of a solution after adding a universal indicator is:

pH range	Indicator	Colour
1-2	Strong acid	Red
3-4	Weak acid	Orange/Yellow
5-6	Neutral	Green
7-8	Weak base	Blue
9-14	Strong base	Dark Blue



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# Class 10 Chemistry Practical Lab Cbse

**Annelies Wilder-Smith**



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