

## Periodic Trends

### ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.
  - a) Al B
  - b) S O
  - c) Br Cl
  - d) Na Al
  - e) O F
  - f) Mg Ca
6. Put the following elements in order from smallest to largest atomic radius **and** explain why:  
C, O, Sn, Sr.

### ELECTRONEGATIVITY

7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.
  - a) Ca Ga
  - b) Li O
  - c) Cl S
  - d) Br As
  - e) Ba Sr
  - f) O S

### GENERAL QUESTIONS

14. Which group tends to form +1 ions? \_\_\_\_\_
15. Which group tends to form +2 ions? \_\_\_\_\_
16. Which group tends to form -1 ions? \_\_\_\_\_
17. Which group tends not to form ions or react? \_\_\_\_\_

# Periodic Table Chapter Test A Answer Key

**AN Whitehead**



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## **Table of Contents Periodic Table Chapter Test A Answer Key**

1. Understanding the eBook Periodic Table Chapter Test A Answer Key
  - The Rise of Digital Reading Periodic Table Chapter Test A Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Chapter Test A Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Chapter Test A Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Chapter Test A Answer Key
  - Personalized Recommendations
  - Periodic Table Chapter Test A Answer Key User Reviews and Ratings
  - Periodic Table Chapter Test A Answer Key and Bestseller Lists
5. Accessing Periodic Table Chapter Test A Answer Key Free and Paid eBooks

- Periodic Table Chapter Test A Answer Key Public Domain eBooks
  - Periodic Table Chapter Test A Answer Key eBook Subscription Services
  - Periodic Table Chapter Test A Answer Key Budget-Friendly Options
6. Navigating Periodic Table Chapter Test A Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Periodic Table Chapter Test A Answer Key Compatibility with Devices
  - Periodic Table Chapter Test A Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Table Chapter Test A Answer Key
  - Highlighting and Note-Taking Periodic Table Chapter Test A Answer Key
  - Interactive Elements Periodic Table Chapter Test A Answer Key
8. Staying Engaged with Periodic Table Chapter Test A Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Table Chapter Test A Answer Key
9. Balancing eBooks and Physical Books Periodic Table Chapter Test A Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Table Chapter Test A Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Chapter Test A Answer Key
- Setting Reading Goals Periodic Table Chapter Test A Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Chapter Test A Answer Key
- Fact-Checking eBook Content of Periodic Table Chapter Test A Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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