

### 3. Atoms elements and compounds



Atomic mass = 23

Proton number = 11

Neutron number = 12 (23 - 11)

mass of electron =  $\frac{1}{1836}$

**ISOTOPE**: atoms of the same element with the same proton number but a different neutron number

- Have the same properties due to the same number of electrons on the outer shell

**Radioisotope**

- testing for leaks in pipes
- treating cancer

#### IONIC BONDING

**Bonding**: oppositely charged ions are strongly attracted by electrostatic force + involves electron transfer

metal + non metal

- metals give electrons to non-metals

**Properties**: strong bonds due to attractive forces

- conduct electricity
- high mp/bp
- charged particles
- soluble in water

#### Macromolecular compounds

**SILICON (IV) OXIDE**

**Properties**



#### COVALENT BONDING

**Bonding**: nuclei of atoms and the shared pair of electrons are strongly attracted by electrostatic forces

non metal + non metal

**Properties**: weak bonds = low mp/bp

- insoluble
- don't conduct

#### 3. Bonding



... the strong ... of ... ..  
 ... ..  
 ... ..

A large box containing a list of properties and examples of covalent bonding. The text is partially illegible but appears to include terms like 'simple molecular', 'giant covalent', and 'intermolecular forces'.

# Igcse Complete Chemistry Notes

**Arturo Cuomo**



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