

**Chapter 12: Atoms**  
**Weightage of marks- 5 Marks**

**I. Short Answer Questions**

1. **\*\*Describe plum pudding model of atom. Who proposed this model?**

➤ JJ Thomson proposed the plum pudding model of atom.

According to this model, the positive charge is uniformly distributed throughout the volume of the atom and the negatively charged electrons are embedded on it.

2. **\*\*Mention one merit and demerit of pudding model of atom.**

➤ **Merit:** It could explain neutral nature of atoms.

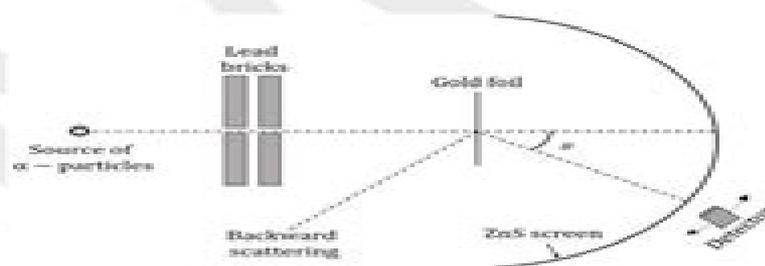
**Demerit:** It could not explain alpha-ray scattering.

**Topics: Alpha particle scattering: Schematic diagram of Geiger-Marsden experiment, observations and conclusion -**

3. **\*\*\*Draw the diagram representing the schematic arrangement of Geiger-Marsden experimental setup for the alpha particle scattering. (2 Marks) - March 2018**

**\*\*\*Explain experimental setup**

- A radioactive source emits  $\alpha$ -particles.
- $\alpha$ -particles are passed through a lead slit to obtain a narrow beam.
- The narrow beam is allowed to fall on a thin gold foil.
- A rotatable detector is used to observe the scattered  $\alpha$ -particles at different angles.
- Detector consists of ZnS screen and a microscope. On striking the screen, alpha particles produce bright flashes
- $\theta$  is the scattering angle.



4. **\*\*\*What were the observations and conclusions of alpha particle scattering?**

➤ **Observation of Geiger-Marsden experiment:**

- Most of the alpha particles pass through the gold foil without any deviation.
- Only about 0.14% of the incident  $\alpha$ -particles scatter by more than  $1^\circ$
- About 1 in 8000 deflect by more than  $90^\circ$ .

**Conclusion:**

- Most of the space in an atom must be empty, as most of the alpha particles pass straight through the foil without any deviation.
- Most of the mass of the atom and its positive charge is concentrated tightly at its centre. This is because, alpha particles carry positive charge and it is 7000 times more massive than an electron. A large repulsive force is required to deflect an alpha particle traveling with high speed, by more than  $90^\circ$ . This is possible only if its mass and positive charge were concentrated at its centre.

**Topics: Rutherford's model of an atom - Derivation of total energy of electron in hydrogen atom in terms of orbit radius.**

5. **\*\*\*Explain Rutherford's model of an atom (planetary model)**

➤ According to Rutherford:

- An atom consists of positively charged core called nucleus.

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**SJ Ball**



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