

Name: _____

Molecule Polarity PhET Lab

A study of electronegativity, bond polarity, and molecular polarity

Introduction:

In this atomic-level simulation, you will investigate how atoms' **electronegativity** value affects the bonds they produce. When two atoms bond, a pair of electrons is shared between atoms. Electronegativity is a measure of a single atom's ability to hoard electrons shared in that bond. In this lab you will work diligently, at your own pace, to answer a number of questions.

To begin, from what you've already learned about the protons and electrons in an atom, what would cause an atom to have a **high** electronegativity value?

Why might an atom have a **low** electronegativity value?

Procedure: **Two Atoms**

- ❖ Turn on (check) all view options. **GRAB** and spin the molecule with your mouse.
- ❖ Take your time and investigate how the binary compound's bond behaves when the atom's electronegativity and orientation are changed. Do not rush through this step.

Describe the bond formed between two atoms with **similar, low** electronegativities in terms of bond character.

Describe the bond formed between two atoms with **similar, high** electronegativities in terms of bond character.

Describe the bond formed between two atoms with **very different** electronegativities in terms of bond character.

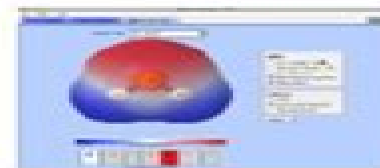
Describe (in your own words) what is meant by **partial charges**, δ^- and δ^+ .

δ^- represents: _____

δ^+ represents: _____

What is **electron density**? _____

How does the density around a partial positive compare to the density of a partial negative?



Molecule Polarity

View

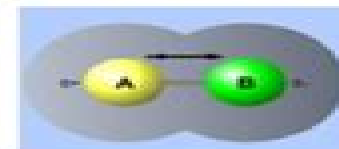
- Bond Dipole \longleftrightarrow
- Partial Charges
- Bond Character

Surface

- none
- Electrostatic Potential
- Electron Density

Electric Field

- on off



Phet Molecular Structure And Polarity Lab Answers

Rachel Sandford



Phet Molecular Structure And Polarity Lab Answers:

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[The Pedersen Memorial Issue](#) R.M. Izatt, J.S. Bradshaw, 2012-12-06 Foreword Charles J Pedersen 1904 1989 Nobel Laureate in Chemistry 1987 This issue is dedicated to the memory of the late Charles J Pedersen in recognition of his outstanding contribution to scientific research culminating in his discovery of crown ethers and their remarkable cation complexing properties and his receipt of the 1987 Nobel Prize in Chemistry Charlie s origin and early years in Korea did not portend the creative work in chemistry which would characterize his later life However we can see in his early years the influence of his Norwegian father and Japanese mother who considered his formal education to be of utmost importance At the age of eight he was sent abroad to Japan for schooling first at a convent school in Nagasaki and two years later at a French American preparatory school in Yokohama run by a Marianist order of Catholic priests and brothers The latter group encouraged him to attend the order s University of Dayton in Ohio where he received a bachelors degree in chemical

engineering Charlie's academic experiences his employment with du Pont and the creative spark which he manifested at an early stage of his scientific career are detailed in the paper in this issue by Herman Schroeder. Schroeder had a long time association with Charlie at du Pont as a co worker supervisor and friend. His recollections provide insight into Charlie's creative mind. In addition they make it clear that a long period of creative work preceded the accidental discovery of the first synthetic crown ether. It is important to note that Charlie's mind was well prepared to recognize the importance of his discovery. The field of macrocyclic chemistry to a large degree had its beginnings with Charlie's discovery. A first person account of his discovery is given as the first paper in this issue. This account was prepared by him and was read at the 12th Symposium on Macrocyclic Chemistry in Hiroshima Japan in 1987 by Herman Schroeder. The growth of this field since Charlie's first publication on the subject in 1967 has been enormous. This growth is evidenced in one segment of the field by the three fold increase in the number of references in two Chemical Reviews articles on thermodynamic quantities associated with cation macrocycle interaction authored by us in 1985 and 1991. Charlie lived to see much of this growth. He saw many of his own predictions of possible uses of crown ethers and related macrocycles realized. Recognition for Charlie came late in his career. He found it satisfying to see so many capable scientists go in so many directions as they applied his discovery to a wide range of chemical and other fields. He made seminal contributions to the broad area known today as molecular recognition. His work illustrates how one individual can make an enormous difference in science. The effect of his life and work on those of us who contributed papers for this issue and on many others is appreciated and is acknowledged by several of the authors in their individual papers. It is entirely appropriate to honor his memory with this special issue.

R M Izatt J S Bradshaw Department of Chemistry Brigham Young University Provo UT 84602 U S A Reprinted from *Journal of Inclusion Phenomena and Molecular Recognition in Chemistry* Volume 12 Nos 1-4 1992

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